

[Problems]

Walker, Physics, 3rd Edition

Chapter 31

Conceptual Questions

(Answers to odd-numbered Conceptual Questions can be found in the back of the book, beginning on page ANS-xx.)

1. Give a reason why the Thomson plum-pudding model does not agree with experimental observations.
2. Give a reason why the Rutherford solar-system model does not agree with experimental observations.
3. Cite one example of how the Bohr model disagrees with the quantum mechanical model of the hydrogen atom.
4. What observation led Rutherford to propose that atoms have a small nucleus containing most of the atom's mass?
5. Do you expect the light given off by **(a)** a neon sign or **(b)** an incandescent lightbulb to be continuous in distribution or in the form of a line spectrum? Explain.
6. In principle, how many spectral lines are there in any given series of hydrogen? Explain.
7. Is there an upper limit to the radius of an allowed Bohr orbit? Explain.
8. **(a)** Is there an upper limit to the wavelength of lines in the spectrum of hydrogen? Explain. **(b)** Is there a lower limit? Explain.
9. The principal quantum number, n , can increase without limit in the hydrogen atom. Does this mean that the energy of the hydrogen atom also can increase without limit? Explain.
10. For each of the following configurations of outermost electrons, state whether the configuration is allowed by the rules of quantum mechanics. If the configuration is not allowed, give the rule or rules that are violated. **(a)** $2d^5$, **(b)** $1p^7$, **(c)** $3p^5$, **(d)** $4g^6$.
11. **(a)** In the quantum mechanical model of the hydrogen atom, there is one value of n for which the angular momentum of the electron must be zero. What is this value of n ? **(b)** Can the angular momentum of the electron be zero in states with other values of n ? Explain.
12. Would you expect characteristic X-rays to be emitted by **(a)** helium atoms or **(b)** lithium atoms in their ground state? Explain.

Physics Ch 31 Packet Answers

RJ Alexander



Physics Ch 31 Packet Answers:

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