

CHAPTER 14 REVIEW*Acids and Bases***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Name the following compounds as acids:

_____ sulfuric acid	a. H_2SO_4
_____ sulfurous acid	b. H_2SO_3
_____ hydrosulfuric acid	c. H_2S
_____ perchloric acid	d. HClO_4
_____ hydrocyanic acid	e. hydrogen cyanide

2. _____ H_2S Which (if any) of the acids mentioned in item 1 are binary acids?

3. Write formulas for the following acids:

_____ HNO_2	a. nitrous acid
_____ HBr	b. hydrobromic acid
_____ H_3PO_4	c. phosphoric acid
_____ CH_3COOH	d. acetic acid
_____ HClO	e. hypochlorous acid

4. Calcium selenate has the formula CaSeO_4 .

_____ H_2SeO_4	a. What is the formula for selenic acid?
_____ H_2SeO_3	b. What is the formula for selenous acid?

5. Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.

6. Write balanced chemical equations for the following reactions of acids and bases:

- a. aluminum metal with dilute nitric acid



- b. calcium hydroxide solution with acetic acid



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