

Name _____

KEY

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REVIEW QUESTIONS FOR FINAL EXAM (2nd Semester)

SECTION 12.1

Heat Transfer: $Q = mC\Delta T$ or $Q = mC(T_f - T_i)$

Specific Heat of Common Substances			
Material	Specific heat J/kg · K	Material	Specific heat J/kg · K
aluminum	900	lead	130
brass	376	nickel	210
carbon	712	silver	235
copper	385	steel	200
glass	840	water	4180
ice	2090	iron	450
oil	2000		

1) Two solid objects are in contact as shown below.



Heat will always flow from object A to object B if object B has a lower

- a) mass b) total energy c) specific heat
d) temperature e) none of the answers

For questions 2-3: The diagram below shows two solids with their temperatures (A=50°C, B=80°C).



2) What occurs when the two solids are placed in contact with each other as shown?

- a) Heat energy flows from A to B. Solid A cools down.
 b) Heat energy flows from A to B. Solid A gets hotter.
c) Heat energy flows from B to A. Solid B cools down.
 d) Heat energy flows from B to A. Solid B gets hotter.

3) The net flow of heat energy will stop from one object to the other when they both have the same

- a) energy **b) temperature**
 c) specific heat d) all the answers are correct

4) How much heat is transferred when a 0.2 kg piece of copper's temperature is raised from 50°C to 80°C?

$$Q = mC\Delta T$$

$$Q = (0.2 \text{ kg})(385 \text{ J/kg} \cdot ^\circ\text{C})(80^\circ\text{C} - 50^\circ\text{C}) = 2310 \text{ J}$$

5) A 0.5 kg block of lead cools down from 30°C to 20°C. How much heat was transferred?

$$Q = (0.5 \text{ kg})(130 \text{ J/kg} \cdot ^\circ\text{C})(20^\circ\text{C} - 30^\circ\text{C})$$

$$Q = -650 \text{ J}$$

6) The heat transfer of a substance is 9400 J. The mass of this substance is 2 kg and its temperature rises from 80°C to 100°C. Based on the specific heat that you will calculate for this substance, what is it?

$$C = \frac{Q}{m\Delta T} = \frac{9400 \text{ J}}{(2 \text{ kg})(100^\circ\text{C} - 80^\circ\text{C})} = 235 \text{ J/kg} \cdot ^\circ\text{C}$$

SILVER

7) From the choices below, which material would be better for a pan handle?

- a) brass 376 b) silver 235 c) copper 385
 d) iron 450 e) zinc 388 f) lead 130

8) Provide a brief description of absolute zero and their temperatures in Kelvin (K) and degrees Celsius (°C).

The lowest temperature when all particle motion stops. $T_C = -273^\circ\text{C}$, $T_K = 0 \text{ K}$.

9) The boiling point of water is 100°C. The boiling point of water in Kelvin (K) will be

$$T_K = 273 + T_C = 273 + 100 = 373 \text{ K}$$

10) A 1-liter of water taken out of a 5-gallon bucket full of water at 20°C will

- a) register a lower temperature than that of the water left in the bucket
b) have less thermal energy than that of the water left in the bucket
 c) have less specific heat than that of the water left in the bucket

11) 243 K in degrees Celsius (°C) is

$$T_C = T_K - 273 = 243 \text{ K} - 273 = -30^\circ\text{C}$$

12) The kinetic molecular theory can explain the existence of thermal energy in a substance.

- a) true** b) false

13) When using a thermometer for measuring the temperature of a substance, one should record the temperature

- a) about 2 minutes after the thermometer has been placed in the substance. That is usually the time needed for the liquid inside to expand or shrink completely.
b) until thermal equilibrium has been reached
 c) until all the heat from the substance has been transferred to the thermometer
 d) until the thermal energy from the substance is equal to the thermal energy of the thermometer
 e) all of the answers

14) A calorimeter contains 1 kg of water at 20°C. A 0.20 kg block of zinc at 120°C is placed in the water. What is the approximate final temperature of the system?

$$T_f = \frac{m_A C_A T_{Ai} + m_B C_B T_{Bi}}{m_A C_A + m_B C_B}$$

$$T_f = \frac{(1)(4180)(20^\circ\text{C}) + (0.2)(388)(120^\circ\text{C})}{(1)(4180) + (0.2)(388)}$$

$$T_f = \frac{83600 + 9312}{4257.6} = 21.8^\circ\text{C}$$

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Table of Contents Physics 1 Final Exam Review Conceptual Questions

1. Understanding the eBook Physics 1 Final Exam Review Conceptual Questions
 - The Rise of Digital Reading Physics 1 Final Exam Review Conceptual Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Physics 1 Final Exam Review Conceptual Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physics 1 Final Exam Review Conceptual Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physics 1 Final Exam Review Conceptual Questions

- Personalized Recommendations
 - Physics 1 Final Exam Review Conceptual Questions User Reviews and Ratings
 - Physics 1 Final Exam Review Conceptual Questions and Bestseller Lists
5. Accessing Physics 1 Final Exam Review Conceptual Questions Free and Paid eBooks
 - Physics 1 Final Exam Review Conceptual Questions Public Domain eBooks
 - Physics 1 Final Exam Review Conceptual Questions eBook Subscription Services
 - Physics 1 Final Exam Review Conceptual Questions Budget-Friendly Options
 6. Navigating Physics 1 Final Exam Review Conceptual Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Physics 1 Final Exam Review Conceptual Questions Compatibility with Devices
 - Physics 1 Final Exam Review Conceptual Questions Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physics 1 Final Exam Review Conceptual Questions
 - Highlighting and Note-Taking Physics 1 Final Exam Review Conceptual Questions
 - Interactive Elements Physics 1 Final Exam Review Conceptual Questions
 8. Staying Engaged with Physics 1 Final Exam Review Conceptual Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physics 1 Final Exam Review Conceptual Questions
 9. Balancing eBooks and Physical Books Physics 1 Final Exam Review Conceptual Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physics 1 Final Exam Review Conceptual Questions
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Physics 1 Final Exam Review Conceptual Questions
 - Setting Reading Goals Physics 1 Final Exam Review Conceptual Questions
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Physics 1 Final Exam Review Conceptual Questions

- Fact-Checking eBook Content of Physics 1 Final Exam Review Conceptual Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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