

CVA Physical Science Unit 6
A6.01.2 Balancing Chemical Equations Assignment



Student Name:

Teacher Name:

Why are we concerned about the amount of each substance present in a chemical reaction? Remember that matter cannot be created nor destroyed during a chemical reaction. This statement is known as the law of conservation of mass.

Chemists use **coefficients** to satisfy the law of conservation of mass. They add coefficients to reactants and products in a chemical equation until there is the same number of each kind of atom on both sides of the equation.

A coefficient in front of a formula indicates the number of molecules of that substance that take part in the chemical reaction. In the reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, 2 molecules of hydrogen must combine with 1 molecule of oxygen to produce 2 molecules of water. When no coefficient appears before a formula, assume that 1 molecule (or atom) of that substance is present. See Lesson 1 in Module 6 for more information on balancing chemical equations.



Balance the following Chemical Equations. Type the correct **coefficient** in front of the reactants and products so you have the same number of atoms of each element on both sides of the equation. Remember to find the number of atoms, simply multiply the coefficient times the subscript for that atom.

Examples:

Fe is an atom of Iron **2Fe** is 2 atoms of Iron

O₂ is a molecule of Oxygen (2 Oxygen atoms bonded together) **2O₂** is 2 Oxygen molecules consisting of 4 Oxygen atoms

(SO₄) is one Sulfate ion (consisting of 1 Sulfur atom and 4 Oxygen atoms), **3(SO₄)** is 3 Sulfate ions (consisting of 3 Sulfur atoms and 12 Oxygen atoms)

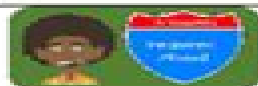
Remember, you can only change the coefficient, not the subscript to balance the equation! Save and submit this assignment. Total points 100 Balancing equation 11 is worth 10 extra credit points!

Example:

Given - $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

Answer: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

Place your cursor in front of the atom or molecule and type in the coefficient that will balance the chemical equations below:



Place coefficients, where needed in the reactants and or products to balance the following chemical reactions. Remember, you must have the same number of atoms of each element on both sides of the chemical equation! (10 pts. each)

1.



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JS Bruner



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