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- Ⓐ Which one of the following decreases in a group from top to bottom?

51. Metallic nature

D) Electronegativity

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- Diagram illustrating the molecular orbital energy levels for the reaction of two nitric oxide (NO) molecules to form dinitrogen dioxide (N₂O₂). The diagram shows the combination of atomic orbitals from two NO molecules to form molecular orbitals in N₂O₂. The energy levels are labeled with their occupancy (number of electrons). The total energy of the products is lower than the total energy of the reactants, indicating an exothermic reaction.

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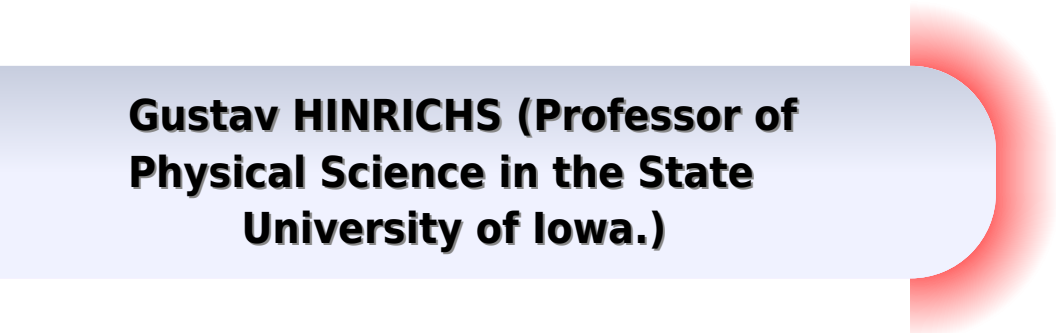
3 - 2 = 6 M

10. Pose any two questions to understand the difference between valency and valency electrons.

11. Explain why bonding angle (HOH) in water is $104^{\circ} 31'$ instead of $109^{\circ} 28'$.

Physical Science Grade 1paper 1

**Gustav HINRICHS (Professor of
Physical Science in the State
University of Iowa.)**



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