

Name key

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. When an atom loses an electron it becomes a positive ion (cation), and when it gains an electron it becomes a negative ion (anion).
2. An element in the halogen group has 7 valence electrons, and wants to gain 1 more to have 8. This is because of the octet rule.
3. An ionic bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal.
4. In a covalent bond, atoms share electrons, whereas in an ionic bond, atoms transfer electrons.
5. (Honors only) When electrons are shared unequally in a covalent bond it is polar, and when they're shared equally it is nonpolar. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the electronegativities.
6. An atom can have up to 8 valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called noble gases.
7. Atoms can gain, lose, or share valence electrons to become more stable.

B. Formulas for Ionic compounds - Fill in the chart below

	Magnesium	Fluorine	Nitrogen	Lithium
# OF VALENCE ELECTRONS	<u>2</u>	<u>7</u>	<u>5</u>	<u>1</u>
OXIDATION #	<u>+2</u>	<u>-1</u>	<u>-3</u>	<u>+1</u>

2. What is the **chemical formula** for each of the following ionic compounds?
(*hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.*)

- a) Potassium and Chlorine - KCl
- b) Calcium and Bromine - CaBr₂
- c) Magnesium and oxygen - MgO

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G Orfield



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United States Armed Forces Institute Catalog United States Armed Forces Institute, 1963 Resources in Education, 1995-04 **Science Curriculum Topic Study** Page Keeley, Joyce Tugel, 2019-09-11 Today's science standards reflect a new vision of teaching and learning How to make this vision happen Scientific literacy for all students requires a deep understanding of the three dimensions of science education disciplinary content scientific and engineering practices and crosscutting concepts If you actively engage students in using and applying these three dimensions within curricular topics they will develop a scientifically based and coherent view of the natural and designed world The latest edition of this best seller newly mapped to the Framework for K-12 Science Education and the Next Generation Science Standards NGSS and updated with new standards and research based resources will help science educators make the shifts needed to reflect current practices in curriculum instruction and assessment The methodical study process described in this book will help readers intertwine content practices and crosscutting concepts The book includes An increased emphasis on STEM including topics in science technology and engineering 103 separate curriculum topic study guides arranged in six categories Connections to content knowledge curricular and instructional implications concepts and specific ideas research on student

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