

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance
 - Volume of an object? ruler or water displacement method
 - Length of an object? Ruler
 - Density of an object? graduated cylinder
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 33 g and has a volume of 3 cm³? 11g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

Physical Science Semester 2 Exam Study Guide

J Elliott



Physical Science Semester 2 Exam Study Guide:

Physical Science Semester 2 Exam Study Guide Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has become more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such is the essence of the book **Physical Science Semester 2 Exam Study Guide**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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Table of Contents Physical Science Semester 2 Exam Study Guide

1. Understanding the eBook Physical Science Semester 2 Exam Study Guide
 - The Rise of Digital Reading Physical Science Semester 2 Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Semester 2 Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Physical Science Semester 2 Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Semester 2 Exam Study Guide
 - Personalized Recommendations
 - Physical Science Semester 2 Exam Study Guide User Reviews and Ratings

- Physical Science Semester 2 Exam Study Guide and Bestseller Lists
- 5. Accessing Physical Science Semester 2 Exam Study Guide Free and Paid eBooks
 - Physical Science Semester 2 Exam Study Guide Public Domain eBooks
 - Physical Science Semester 2 Exam Study Guide eBook Subscription Services
 - Physical Science Semester 2 Exam Study Guide Budget-Friendly Options
- 6. Navigating Physical Science Semester 2 Exam Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Semester 2 Exam Study Guide Compatibility with Devices
 - Physical Science Semester 2 Exam Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Semester 2 Exam Study Guide
 - Highlighting and Note-Taking Physical Science Semester 2 Exam Study Guide
 - Interactive Elements Physical Science Semester 2 Exam Study Guide
- 8. Staying Engaged with Physical Science Semester 2 Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Semester 2 Exam Study Guide
- 9. Balancing eBooks and Physical Books Physical Science Semester 2 Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Semester 2 Exam Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Semester 2 Exam Study Guide
 - Setting Reading Goals Physical Science Semester 2 Exam Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Semester 2 Exam Study Guide
 - Fact-Checking eBook Content of Physical Science Semester 2 Exam Study Guide
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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