

QUESTION 1: MULTIPLE-CHOICE QUESTIONS

Various options are provided as possible answers to the following questions. Choose the answer and write only the letter (A-D) next to the question number (1.1 - 1.10) in the ANSWER BOOK, for example 1.5 D.

1.1 Which ONE of the following is a molecule with a multiple bond?

- A N_2
- B NH_3
- C OF_2
- D HOCl

(2)

1.2 The molecular formula of ethyne is C_2H_2 .

The shape of the ethyne molecule as predicted by the VSEPR theory is:

- A Trigonal planar.
- B Octahedral.
- C Linear.
- D Bent.

(2)

1.3 The table below indicates the boiling points of four liquids.

SUBSTANCE	BOILING POINT ($^{\circ}\text{C}$)
Water	100
Methylated spirits	78,5
Ethanol	78,37
Acetone	56

Which ONE of the above liquids has the lowest surface tension?

- A Water.
- B Methylated spirits.
- C Ethanol.
- D Acetone.

(2)

1.4 The predominant forces between the molecules in an ice crystal are called . . .

- A London forces.
- B Hydrogen bonding.
- C Induced dipole forces.
- D Polar covalent bonding.

(2)
[8]

Physical Science Past Exam Papers 2013

RC Schank



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