



NATIONAL
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2
(CHEMISTRY)
EXEMPLAR

MARKS: 100
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.
Use blue or black ink for writing. Do not use red ink.

QUESTION 3

PHYSICAL SCIENCES PAPER 2

INTERMOLECULAR FORCES

GRADE 11

NOVEMBER 2020

Examiner's use: _____ Physical Sciences P2

QUESTION 3 starts on a NEW page.

- 3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the ΔH_{vap} and ΔH_{fusion} for 1-hexanol in kJ mol⁻¹ and 1 mol of 1-hexanol.

Compound	Molecular mass (M _r)	Boiling point (°C)
A	102	16,2
B	102	21

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol composed of one (A) or two (B) and list one of its intermolecular forces? (2)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and its boiling point of these molecules. (2)
- 3.1.4 What (one) of the compounds (A or B) has a (threefold) greater mass (molar mass)? (2)
- 3.1.5 Which compound (A or B) will have the higher vapour pressure? (2)
- 3.1.6 Which compound (A or B) will have the higher surface tension? (2)

Compound B (M_r 102) is shown in the table. Use a suitable intermolecular force that both compounds A and B have to identify (one) boiling point of 16,2 °C.

- 3.1.6 Explain why the boiling point of compound B is (threefold) higher than compound A and B is (threefold) higher than 1-hexanol in terms of intermolecular forces. (2)

- 3.2 Substances C1 is added to water whereas 1-pentanol (substance C2) is not added. It separates out when the water is shaken. (2)



- 3.2.1 Is substance C1 (A or B), and the water (A or B)? (2)
- 3.2.2 Explain the process in QUESTION 3.2.1 about the solubility in the intermolecular forces involved. (2)

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Page 10 of 10

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Scott C. Dulebohn



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Table of Contents Physical Science Paper2 November 2014

1. Understanding the eBook Physical Science Paper2 November 2014
 - The Rise of Digital Reading Physical Science Paper2 November 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Paper2 November 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Paper2 November 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Paper2 November 2014
 - Personalized Recommendations
 - Physical Science Paper2 November 2014 User Reviews and Ratings
 - Physical Science Paper2 November 2014 and Bestseller Lists

5. Accessing Physical Science Paper2 November 2014 Free and Paid eBooks
 - Physical Science Paper2 November 2014 Public Domain eBooks
 - Physical Science Paper2 November 2014 eBook Subscription Services
 - Physical Science Paper2 November 2014 Budget-Friendly Options
6. Navigating Physical Science Paper2 November 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Paper2 November 2014 Compatibility with Devices
 - Physical Science Paper2 November 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Paper2 November 2014
 - Highlighting and Note-Taking Physical Science Paper2 November 2014
 - Interactive Elements Physical Science Paper2 November 2014
8. Staying Engaged with Physical Science Paper2 November 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Paper2 November 2014
9. Balancing eBooks and Physical Books Physical Science Paper2 November 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Paper2 November 2014
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Paper2 November 2014
 - Setting Reading Goals Physical Science Paper2 November 2014
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Paper2 November 2014
 - Fact-Checking eBook Content of Physical Science Paper2 November 2014
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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