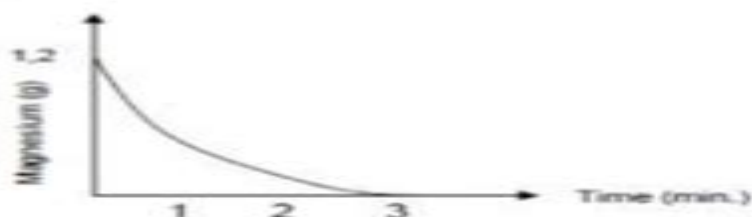


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records the change in the mass of magnesium as the reaction proceeds at regular intervals. The balanced chemical equation for the reaction is:



The change in the mass of magnesium during the reaction is shown on the graph below.



6.2.1 Identify the limiting agent in this reaction. Give a reason for the answer. (2)

6.2.2 Calculate the number of moles of unreacted hydrochloric acid in the test tube after 3 minutes. (7)

[19]

Hide Answers

QUESTION/VRAAG 6

6.1.1 $2\text{H}_2\text{O}_2\text{(aq)} \checkmark \rightarrow 2\text{H}_2\text{O(l)} + \text{O}_2\text{(g)} \checkmark$ (2)

6.1.2 The catalyst lowers the activation energy of the reaction ✓✓
 Accept: catalyst speeds up the reaction ✓✓
 'n Katalisator verlaag die aktiveringsenergie van die reaksie ✓✓
 Aanvaar: katalisator laat die reaksie vinniger plaasvind (2)

6.1.3 **OPTION 1/OPSIE 1**
 $n = \frac{V}{V_m} \checkmark$
 $n = \frac{0,6}{24,45} \checkmark$
 $n = 0,0245 \text{ mole/mol O}_2 \text{ produced/gevorm} \checkmark$
 $\text{H}_2\text{O}_2 : \text{O}_2$
 $2 : 1 \checkmark$
 $n = 0,049 \text{ mole/mol H}_2\text{O}_2 \text{ reacted/reageer} \checkmark$
 $n = \frac{m}{M} \checkmark$
 $0,049 = \frac{m}{34} \checkmark$

OPTION 2/OPSIE 2
 From the balanced equation:
 Vanaf gebalanseerde vergelyking:
 $68\text{g H}_2\text{O}_2 \rightarrow 24,45 \text{ dm}^3 \text{ O}_2 \checkmark \checkmark$
 $\text{X g H}_2\text{O}_2 \rightarrow 600 \times 10^{-3} \text{ dm}^3 \checkmark$
 $\text{X} = \frac{68 \times 0,6}{24,45} \checkmark$
 $\text{X} = 1,67 \text{ g} \checkmark$

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Ian Pickup



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