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Unit 7: Balancing Chemical Reactions · Worksheet 2

Balance the following equations by inserting the proper coefficients.

- 1. $_2$ _SO₂ + $_1$ _O₂ \rightarrow $_2$ _SO₃
- 2. $2 \text{ CH4} + 3 \text{ O2} \rightarrow 2 \text{ CO} + 4 \text{ H2O}$
- 2 P + 3 G2 → 2 PG3
- 2_CO + 1_O₂ → 2_CO₂
- 5. $_2$ CH4 + $_1$ O2 \rightarrow $_2$ CH3OH
- 6. $_2$ _Li + $_1$ _Br₂ \rightarrow $_2$ LiBr
- 7. $_2AI_2O_3 \rightarrow _4AI + _3O_2$
- 8. $_{2}$ Na + $_{2}$ H₂O \rightarrow $_{2}$ NaOH + $_{1}$ H₂
- 9. $_{6}$ _CO₂ + $_{6}$ _H₂O \rightarrow $_{1}$ _C₆H₁₂O₆ + $_{6}$ _O₂
- 10. _1_H2SO4 + _2_NaCl → _2_HCl + _1_Na2SO4

Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:

- 11. 3 H₂ + 1 SO₂ → 1 H₂S + 2 H₂O
- 12. $_2$ CaCO3 + $_2$ SO2 + $_1$ O2 \rightarrow $_2$ CaSO4 + $_2$ CO2
- 13. 2_AgNO₃ + 1_CaCl₂ → 2_AgCl + 1_Ca(NO₃)₂
- 14. 2 HCl + 1 Ba(OH)2 → 1 BaCl2 + 2 H2O
- 15. _1_H3PO4 + _3_NaOH → _1_Na3PO4 + _3_H2O
- 16. _1_Pb(NO3)2 + _2_KI → _1_PbI2 + _2_KNO3
- 17. $3 \text{ O}_1\text{ O} + 2 \text{ NH}_3 \rightarrow 1 \text{ N}_2 + 3 \text{ O}_1 + 3 \text{ H}_2\text{ O}$
- 18. $_1$ C₂H₅OH + $_3$ C₂ \rightarrow $_2$ CO₂ + $_3$ H₂O
- 19. $_2$ C2H6 + $_3$ O2 \rightarrow $_2$ CH3COOH + $_2$ H2O
- 20. $3NO_2 + 1H_2O \rightarrow 2HNO_3 + 1NO$

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Thirumalaisamy P. Velavancorresponding

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