

on
sides again
10/12

Date _____ Name _____
Section _____ Time _____
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Pre-Lab Study Questions

1. Why are burning candles and rusting nails examples of chemical change?

Rust is oxidation

Candles are combustion
things undergo one of the 4 criteria
formation of gas, formation of solid, formation
of new substances, or heat produced/absorbed.

2. What is included in a chemical equation?

a reactant and

a product

3. How does a combination reaction differ from a decomposition reaction?

Combination - two elements combine to form
a compound.

Decomposition - the compound splits
apart.

4. Balance each of the following reactions, and identify the type of reaction.

Unbalanced Equation	Type of Reaction
a. $2\text{Al}(s) + \text{Fe}_2\text{O}_3(s) \xrightarrow{\Delta} \text{Al}_2\text{O}_3 + \text{Fe}(l)$	single replacement
b. $2\text{KClO}_3(s) \xrightarrow{\Delta} 2\text{KCl}(s) + 3\text{O}_2(g)$	decomposition
c. $\text{I}_2(s) + \text{Cl}_2(g) \xrightarrow{\Delta} 2\text{ICl}(g)$	combination
d. $\text{C}_2\text{H}_4(g) + \text{O}_2(g) \xrightarrow{\Delta} \text{CO}_2(g) + \text{H}_2\text{O}(g)$	double replacement
e. $2\text{CrCl}_3(aq) + 3\text{H}_2\text{S}(g) \longrightarrow \text{Cr}_2\text{S}_3(s) + 6\text{HCl}(aq)$	double replacement

5. Complete and balance each of the following reactions.

Reactants	Type of Reaction
a. $\text{K}(s) + \text{N}_2(g) \longrightarrow 2\text{KN}(g)$	Combination
b. $\text{C}_3\text{H}_{12}(g) + \text{O}_2(g) \xrightarrow{\Delta}$	Combustion
c. $\text{Al}(s) + \text{CuSO}_4(aq) \longrightarrow$	Single Replacement
d. $\text{CoCl}_2(aq) + \text{AgNO}_3(aq) \longrightarrow$	Double Replacement
e. $\text{MgCO}_3(s) \xrightarrow{\Delta} \text{MgO}(s) + \text{CO}_2(g)$	Decomposition

Pigeon Pre Lab Questions

Liyong Dong

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Table of Contents Pigeon Pre Lab Questions

1. Understanding the eBook Pigeon Pre Lab Questions
 - The Rise of Digital Reading Pigeon Pre Lab Questions
 - Advantages of eBooks Over Traditional Books
2. Identifying Pigeon Pre Lab Questions
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pigeon Pre Lab Questions
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pigeon Pre Lab Questions
 - Personalized Recommendations
 - Pigeon Pre Lab Questions User Reviews and Ratings
 - Pigeon Pre Lab Questions and Bestseller Lists

5. Accessing Pigeon Pre Lab Questions Free and Paid eBooks
 - Pigeon Pre Lab Questions Public Domain eBooks
 - Pigeon Pre Lab Questions eBook Subscription Services
 - Pigeon Pre Lab Questions Budget-Friendly Options
6. Navigating Pigeon Pre Lab Questions eBook Formats
 - ePub, PDF, MOBI, and More
 - Pigeon Pre Lab Questions Compatibility with Devices
 - Pigeon Pre Lab Questions Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pigeon Pre Lab Questions
 - Highlighting and Note-Taking Pigeon Pre Lab Questions
 - Interactive Elements Pigeon Pre Lab Questions
8. Staying Engaged with Pigeon Pre Lab Questions
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pigeon Pre Lab Questions
9. Balancing eBooks and Physical Books Pigeon Pre Lab Questions
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pigeon Pre Lab Questions
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pigeon Pre Lab Questions
 - Setting Reading Goals Pigeon Pre Lab Questions
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pigeon Pre Lab Questions
 - Fact-Checking eBook Content of Pigeon Pre Lab Questions
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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