

# ATOMIC STRUCTURE AND THE PERIODIC TABLE CHAPTER 4 WORKSHEET

## PART A

Given the following isotopes, determine the atomic number, the mass number, the number of protons, electrons and neutrons.

| Isotope Symbol        | Atomic Number | Mass Number | Protons | Electrons | Neutrons | Isotope Name   |
|-----------------------|---------------|-------------|---------|-----------|----------|----------------|
| $^{131}_{53}\text{I}$ | 53            | 131         | 53      | 53        | 78       | iodine - 131   |
| $^{35}_{16}\text{S}$  | 16            | 35          | 16      | 16        | 19       | sulphur - 35   |
| $^4_2\text{He}$       | 2             | 4           | 2       | 2         | 2        | helium - 4     |
| $^{27}_{13}\text{Al}$ | 13            | 27          | 13      | 13        | 14       | aluminium - 27 |
| $^{81}_{36}\text{Kr}$ | 36            | 81          | 36      | 36        | 45       | krypton - 81   |
| $^{81}_{37}\text{Rb}$ | 37            | 81          | 37      | 37        | 44       | rubidium - 81  |

## PART B

Complete the following chart by writing the symbol for the isotope of the following elements. In addition, give the number of protons, electrons, mass number and atomic number and complete the element name.

| Element Name    | Neutrons | Protons | Electrons | Mass Number | Atomic Number | Isotope Symbol        |
|-----------------|----------|---------|-----------|-------------|---------------|-----------------------|
| Uranium-<br>237 | 145      | 92      | 92        | 237         | 92            | $^{237}_{92}\text{U}$ |
| Chlorine-<br>45 | 28       | 17      | 17        | 45          | 17            | $^{45}_{17}\text{Cl}$ |
| Oxygen-<br>17   | 9        | 8       | 8         | 17          | 8             | $^{17}_8\text{O}$     |
| Boron-<br>11    | 6        | 5       | 5         | 11          | 5             | $^{11}_5\text{B}$     |
| Beryllium-<br>9 | 5        | 4       | 4         | 9           | 4             | $^9_4\text{Be}$       |
| Hydrogen-<br>2  | 1        | 1       | 1         | 2           | 1             | $^2_1\text{H}$        |
| Carbon-<br>14   | 8        | 6       | 6         | 14          | 6             | $^{14}_6\text{C}$     |

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**K Morrison**



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