The equilibrium constant for the reaction.

$$H_{2(g)} + I_{2(g)} \rightleftharpoons 2HI_{(g)}$$

at 448°C is 50.

a) If 1.0 mol of H2 is mixed with 1.0 mol of I2 in a 0.50 L container and allowed to react at 448°C, what is the equilibrium [HI]? Initial [H2] = 1.000 = 2.0M, [T2] = 2.0M, [HI] = 0

H, + I2 = 2 HI 2.0 2.0 + 2× - X 2.0-x 2.0-X 2.0-1.559 2.0-1.559

$$keq = \frac{EHIJ^{2}}{EH_{2}JEI_{2}J}$$

50. =  $\frac{(2x)^{2}}{(2.0-x)^{2}}$ 

Keq =  $\frac{\text{EHI}_{3}^{2}}{\text{EH}_{2}\text{I}\text{EI}_{2}\text{I}}$  | 7.071 (2-x) = 2x 14.142-7.071x = 2x 14.142 = 9.071 x x =  $\frac{(2x)^{2}}{(2.0-x)^{2}}$  |  $x = \frac{14.142}{9.071} = 1.559$ 

- How many moles of HI are formed at equilibrium? (Actual yield)
- 3.118M x 0.50L = 1.6 mal

Answer 1.6 moles of HI

Given Keq for the reaction:

$$PCl_{5(g)} \rightleftharpoons PCl_{3(g)} + Cl_{2(g)}$$

is 0.042 at 250°C, what will happen if 2.50 mol of PCIs, 0.600 mol of CI2 and 0.600 mol of PCI<sub>3</sub> are placed in a 1.00 flask at 250°C ? (Will the reaction shift left, right, or not occur at all?)

$$Keq = \frac{EPCI_3 3 ECI_2 3}{EPCI_5 3}$$

Trial  $Keq = \frac{(0.600)(0.600)}{(2.50)}$ 

Trial  $Keq = 0.144$ 

Actual  $Keq = 0.042$ 

- since Trial Keg > Keg The reaction will shift to the LEFT as it approaches equilm So[PCIs] will increase while [PCh3 and [Ch3 decrease.

Answer [HI] = 3.1 M

Answer Shift LEFT

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**Padhraic Smyth** 

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