

Nelson Biology VCE Units 3 & 4

Answers to Chapter Questions

Chapter 1

Review questions

- 1 Both non-living and living things are composed of matter. Matter is made up of atoms.
- 2 Living things are capable of reproduction, growth and movement.
- 3 6 electrons
- 4 The proton defines the element.
- 5 The atomic number would be 15 and the mass number would be 32. The symbol for phosphorus is P. The standard notation is $^{32}_{15}\text{P}$.
- 6 The number of neutrons may vary in all the isotopes of a particular element. Hence, an element may have different mass numbers.
- 7 The four major elements found in living things are oxygen, carbon, hydrogen and nitrogen.
- 8 Radioactive isotopes are used as a source of radiation in therapy and as tracers to tag particular compounds.
- 9 A compound is a pure substance made up of more than one type of atom held together by chemical bonds.
- 10 Ionic compounds are composed of cations and anions, which form ionic bonds. Molecular compounds result from atoms sharing their electrons to form covalent bonds. Molecular compounds vary in their solubility in water depending upon their polar nature. Ionic compounds are generally more soluble in water than molecular compounds.
- 11 Molecular compounds result from the sharing of electrons between non-metal elements, whereas ionic compounds are composed of ions and result from the transfer of electrons between metallic and non-metallic elements. As a result of the charges carried, ionic compounds tend to be more readily dissolved in water compared to molecular compounds.
- 12 A chemical bond is the force of attraction holding atoms together in chemical molecules.
- 13 Molecules that have permanent regions of opposite charge are termed 'polar'. Non-polar molecules have no permanent regions of opposite charge.
- 14 Water molecules are highly polar and attract one another, forming hydrogen bonds. Hydrogen bonding determines many of the unique properties of water.
- 15 Hydrophilic molecules, such as alcohol, have regions of opposite charge that readily form hydrogen bonds with water molecules and hence dissolve.

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Table of Contents Nelson Biology Units And Answers Chapter

1. Understanding the eBook Nelson Biology Units And Answers Chapter
 - The Rise of Digital Reading Nelson Biology Units And Answers Chapter
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Biology Units And Answers Chapter
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Biology Units And Answers Chapter
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Biology Units And Answers Chapter
 - Personalized Recommendations
 - Nelson Biology Units And Answers Chapter User Reviews and Ratings
 - Nelson Biology Units And Answers Chapter and Bestseller Lists
5. Accessing Nelson Biology Units And Answers Chapter Free and Paid eBooks

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- Nelson Biology Units And Answers Chapter eBook Subscription Services
- Nelson Biology Units And Answers Chapter Budget-Friendly Options
- 6. Navigating Nelson Biology Units And Answers Chapter eBook Formats
 - ePub, PDF, MOBI, and More
 - Nelson Biology Units And Answers Chapter Compatibility with Devices
 - Nelson Biology Units And Answers Chapter Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nelson Biology Units And Answers Chapter
 - Highlighting and Note-Taking Nelson Biology Units And Answers Chapter
 - Interactive Elements Nelson Biology Units And Answers Chapter
- 8. Staying Engaged with Nelson Biology Units And Answers Chapter
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Biology Units And Answers Chapter
- 9. Balancing eBooks and Physical Books Nelson Biology Units And Answers Chapter
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Biology Units And Answers Chapter
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Nelson Biology Units And Answers Chapter
 - Setting Reading Goals Nelson Biology Units And Answers Chapter
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Nelson Biology Units And Answers Chapter
 - Fact-Checking eBook Content of Nelson Biology Units And Answers Chapter
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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