



# Samsung Captivate Slide Manual

**C Cleary**



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web a similar unit of concentration is molality  $m$  which is defined as the number of moles of solute per kilogram of solvent not per liter of solution molality =  $\frac{\text{moles solute}}{\text{kilograms solvent}}$  mathematical manipulation of molality is the

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web 1 molarity  $M$  = mass molar mass  $\times \frac{0.4500 \text{ L}}{0.825 \text{ g}} \times \frac{141.9579 \text{ g/mol}}{0.0129 \text{ m}^3}$  molality =  $\frac{0.825 \text{ g}}{141.9579 \text{ g/mol}} \times \frac{0.00581158 \text{ mol}}{0.00581158 \text{ mol}} \times \frac{0.4500 \text{ kg}}{0.0129 \text{ m}^3}$  mole fraction =  $\frac{2 \text{ HPO}_4}{0.825 \text{ g}} \times \frac{141.9579 \text{ g/mol}}{0.00581158 \text{ mol}} \times \frac{0.4500 \text{ g}}{18.015 \text{ g/mol}} \times \frac{24.97918401 \text{ mol}}{1}$

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web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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web what mass of  $\text{CoCl}_2$  in grams is needed for the solution express the answer using 3 significant figures

**8 3 concentrations of solutions problems chemistry libretexts** - Dec 13 2022

web the hardness of water hardness count is usually expressed in parts per million by mass of  $\text{CaCO}_3$  which is equivalent to milligrams of  $\text{CaCO}_3$  per liter of water what is the molar concentration of  $\text{Ca}^{2+}$  ions in a water sample with a hardness count of 175 mg  $\text{CaCO}_3$  L

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web answer the mass percentage of an element of a solution is defined as the mass of the solute in grammes present in 100 g of the solution it s written like this begin array l mass of a component frac mass of component in solution total mass of solution x 100 end array

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web aug 13 2017 molarity practice problems this chemistry video tutorial explains how to solve common molarity problems it discusses how to calculate the concentration of a solution given the mass in grams