# SALT ANALYSIS

# 디

## SCHEME OF ANALYSIS

| EXPERIMENT  | OBSERVATION  | INFERENCE   |
|---|--|---|
| PRELIMINARY EXAMINATIONS :  |  | 4   |
| 1) Noted the colour   | Colourless   | Absence of<br>Cu2+,Fe2+,Fe3+,Mn2+,Co2+,Ni2+   |
|   | Blue   | Pre. of Cu <sup>2</sup> *   |
|   | Brown  | Pre. of Fe <sup>3+</sup>  |
|   | Green  | Pre. of Ni <sup>2</sup> *   |
|   | Pale pink  | Pre. of Mn2*  |
|   | Pink   | Pre. of Co2*  |
| Noted the solubility    a) In water    b) In dil.HCl  | a) Soluble and no<br>ppt with Na <sub>2</sub> CO <sub>3</sub><br>b) Insoluble                                | a) Pre. of NH <sub>4</sub> * b) Pre. of Pb <sup>2</sup> *   |
| <ol> <li>Flame Test:         <ul> <li>A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame.</li> </ul> </li> </ol>            | a) Blue surrounded by green b) Brick red c) Crimson red d) Bluish white e) Grassy green on prolonged heating | a) Pre. of Cu <sup>2+</sup> b) Pre. of Ca <sup>2+</sup> c) Pre. of Sr <sup>2+</sup> d) Pre. of Pb <sup>2+</sup> e) Pre. of Ba <sup>2+</sup> |
| <ol> <li>Ash Test:         <ul> <li>A little of the salt is heated with con.HNO<sub>3</sub> and cobalt nitrate solution. A filter paper is dipped in it and burnt.</li> </ul> </li> </ol> | a) Green ash b) Blue tinted ash c) Pale pink ash   | a) Pre. of Zn <sup>2+</sup> b) Pre. of Al <sup>2+</sup> c) Pre. of Mg <sup>2+</sup>   |
| SYSTEMATIC ANALYSIS OF ANIONS   | :  | *   |
| <ol> <li>Dil.H<sub>2</sub>SO<sub>4</sub> test:         To the salt add dil.H<sub>2</sub>SO<sub>4</sub>, warm if required.     </li> </ol>   | a) Evolution of<br>Colourless,<br>odourless gas<br>with brisk<br>effervescence,<br>turns limewater<br>milky. | a) Pre. of carbonates   |
|   | <ul> <li>b) Smell of vinegar</li> </ul>  | b) Pre. of acetate  |

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**Annelies Wilder-Smith** 

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