

## Model 2 – Ternary Acids (Oxyacids)

| Acid                    | Name of Acid in Aqueous Solution | Cation (+)              | Polyatomic Anion (-) | Polyatomic Anion Name |
|-------------------------|----------------------------------|-------------------------|----------------------|-----------------------|
| $\text{HClO}_3$         | Chloric acid                     | $\text{H}_3\text{O}^+$  | $\text{ClO}_3^{-1}$  | Chlorate              |
| $\text{H}_2\text{SO}_3$ | Sulfurous acid                   | $2\text{H}_3\text{O}^+$ | $\text{SO}_3^{-2}$   | Sulfite               |
| $\text{H}_2\text{SO}_4$ | Sulfuric acid                    | $2\text{H}_3\text{O}^+$ | $\text{SO}_4^{-2}$   | Sulfate               |
| $\text{H}_3\text{PO}_3$ | Phosphorous acid                 | $3\text{H}_3\text{O}^+$ | $\text{PO}_3^{-3}$   | Phosphite             |
| $\text{H}_3\text{PO}_4$ | Phosphoric acid                  | $3\text{H}_3\text{O}^+$ | $\text{PO}_4^{-3}$   | Phosphate             |
| $\text{HNO}_3$          | Nitric acid                      | $\text{H}_3\text{O}^+$  | $\text{NO}_3^{-1}$   | Nitrate               |
| $\text{HNO}_2$          | Nitrous acid                     | $\text{H}_3\text{O}^+$  | $\text{NO}_2^{-1}$   | Nitrite               |
| $\text{H}_2\text{CO}_3$ | Carbonic acid                    | $2\text{H}_3\text{O}^+$ | $\text{CO}_3^{-2}$   | Carbonate             |

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

*Ternary acids contain three elements and they also contain the element oxygen.*

b. What number do you think the prefix "ter-" refers to?

*"ter" refers to three*

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming  $\text{HClO}_3$ ?

*$\text{ClO}_3^{-1}$  has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula  $\text{HCl}$ , not  $\text{HClO}_3$ .*

10. Write a rule for naming ternary acids.

*- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.*

*- "ate" is replaced with "ic" and "ite" is replaced with "ous".*

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



# Polyatomic Ions Pogil Answer Key

**RJ Shavelson**



**Polyatomic Ions Pogil Answer Key:**

**The Particulate Nature of Polyatomic Ions ,2007**

## Unveiling the Magic of Words: A Review of "**Polyatomic Ions Pogil Answer Key**"

In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is really awe-inspiring. Enter the realm of "**Polyatomic Ions Pogil Answer Key**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve into the book's central themes, examine its distinctive writing style, and assess its profound affect on the souls of its readers.

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### **Table of Contents Polyatomic Ions Pogil Answer Key**

1. Understanding the eBook Polyatomic Ions Pogil Answer Key
  - The Rise of Digital Reading Polyatomic Ions Pogil Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Polyatomic Ions Pogil Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Polyatomic Ions Pogil Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Polyatomic Ions Pogil Answer Key
  - Personalized Recommendations
  - Polyatomic Ions Pogil Answer Key User Reviews and Ratings
  - Polyatomic Ions Pogil Answer Key and Bestseller Lists

5. Accessing Polyatomic Ions Pogil Answer Key Free and Paid eBooks
  - Polyatomic Ions Pogil Answer Key Public Domain eBooks
  - Polyatomic Ions Pogil Answer Key eBook Subscription Services
  - Polyatomic Ions Pogil Answer Key Budget-Friendly Options
6. Navigating Polyatomic Ions Pogil Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Polyatomic Ions Pogil Answer Key Compatibility with Devices
  - Polyatomic Ions Pogil Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Polyatomic Ions Pogil Answer Key
  - Highlighting and Note-Taking Polyatomic Ions Pogil Answer Key
  - Interactive Elements Polyatomic Ions Pogil Answer Key
8. Staying Engaged with Polyatomic Ions Pogil Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Polyatomic Ions Pogil Answer Key
9. Balancing eBooks and Physical Books Polyatomic Ions Pogil Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Polyatomic Ions Pogil Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Polyatomic Ions Pogil Answer Key
  - Setting Reading Goals Polyatomic Ions Pogil Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Polyatomic Ions Pogil Answer Key
  - Fact-Checking eBook Content of Polyatomic Ions Pogil Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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