

Periodic and Groups: Valence Electrons: Mastery

Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-1, 2A-2, 3A-3, 4A-4, 5A-5, 6A-6, 7A-7, 8A-8.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group					Group	
	1A	2A	3A	4A	5A	6A	7A	8A	1A
1	H	He							Li
2	B	C	N	O	F	Ne			Na
3	Al	Si	P	S	Cl	Ar			K
4	Ga	Ge	As	Se	Br	Kr			Rb

Transition metals are in between these.

Thinking like the game "Monopoly", answer that can describe the period and group.

What element is in Group 1, and Period 10? (Group 1 is in the first column, 10 is the 10th column.)

What element is in Group 10, and Period 1?

What element is in Group 2A, and Period 10?

What group number is in Group 10?

Group: _____ Period: _____

What group number is in Group 10?

Group: _____ Period: _____

Valence Electrons

Valence Electrons are the outermost electrons in an atom. Each group contains the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Group 17



7 valence electrons

Hydrogen and Helium are both in Group 1A-1, but Helium has 2 valence electrons.

Group 16, tellurium 54, and Group 16, selenium 34. How many valence electrons do they have?

How many electrons?

Group 18: - These are noble gases that have a full shell of electrons. Elements are in Group 18 and they are called noble gases. Helium, neon, argon, krypton, xenon, and radon are noble gases. They have 2, 8, 18, 32, 54, and 86 valence electrons respectively. All noble gases have 8 valence electrons.

Periods And Groups Valence Electrons Masses Answers

J Rink



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