

Average Atomic Mass

How can the masses of the particles within an atom be determined?

Objectives

After observing three gases, three ions, and naturally occurring three isotopes, the masses of these particles, ions, and atoms should be determined. Comparison of particles including those existing in the natural elements. How does having different kinds of particles (protons and neutrons) affect the mass of the particles? How does having different kinds of particles affect the mass of the atom? How does having different kinds of particles affect the mass of the element?

Model 1 - A Strip of Magnesium Metal



1. What are the particles making up each Mg atom in Model 1?
They are made of three particles: protons, neutrons, and electrons.
2. What are the mass numbers of the naturally occurring isotopes of magnesium (shown in Model 1)?
24, 25, 26.
3. How all of the masses of magnesium in Model 1 have the same number of protons?
Yes, but the mass numbers of the atoms are different, they contain different numbers of neutrons.
4. How the weights of the naturally occurring magnesium shown in Model 1 relate to the fact that the mass numbers of the three isotopes and the number of atoms of each isotope present.

	^{24}Mg	^{25}Mg	^{26}Mg
Number of atoms	10	5	5

5. What's the mass of magnesium in the atom numbers in Model 1?
 ^{24}Mg
6. What are Model 1 and the data given suggest in Question 4, the mass of 14 atoms of magnesium (approximately twice the mass number of the isotope) will be (about 140) for Model 1.
As a group of 14 magnesium atoms, there will be approximately one ^{24}Mg , one ^{25}Mg , and eight ^{26}Mg atoms.

Periodicity Of Elements Pogil Key

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