

Periodic Trends Worksheet

Using a few minutes before today's activity, complete the following questions. Please post your responses on the board when you are done.

Additional help on the periodic trends can be found on this video: <https://www.youtube.com/watch?v=0D8Wp-2GgY0>

1. What does electronegativity measure and how does it change across the periodic table? (electronegativity is largest) in which direction, electronegativity, and location.
2. Ionization energy measures how a larger atomic radius affects electronegativity.
3. Ionization energy measures how a larger atomic radius affects electronegativity.
 - a. Fluorine is Li
 - b. It is wrong
 - c. Cl is Cl
4. Ionization energy measures how a larger atomic radius affects electronegativity.
 - a. It is Cl
 - b. Calcium is
 - c. Li is Li
5. Ionization energy measures how a larger atomic radius affects electronegativity.
 - a. Cl is Cl
 - b. Fluorine is Li
 - c. It is Br
6. Ionization energy measures how a larger atomic radius affects electronegativity.
 - a. It is Cl
 - b. Fluorine is
 - c. It is Br
7. Ionization energy measures how a larger atomic radius affects electronegativity.
 - a. It is Cl
 - b. Fluorine is
 - c. It is Br
8. Ionization energy measures how a larger atomic radius affects electronegativity.
 - a. It is Cl
 - b. Fluorine is
 - c. It is Br
9. Ionization energy measures how a larger atomic radius affects electronegativity.
 - a. It is Cl
 - b. Fluorine is
 - c. It is Br
10. Chlorine is the most electronegative of the elements in the periodic table.

Periodic Trends Questions And Answers

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