

Annotated periodic table of the elements

1	H 1.01 Hydrogen												13	14	15	16	17	18
2	3	4	Main common oxidation states: electron affinity (kJ/mol) (Mulliken)										5	6	7	8	9	10
11	12	13											19	20	21	22	23	24
37	38	39											55	56	57	58	59	60
87	88	89											103	104	105	106	107	108
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- Group names*
- 1 Hydrogen & the alkali metals
 - 2 Alkaline earth metals
 - 11 Coinage metals
 - 12 Volatile metals
 - 13 Icosagens
 - 14 Chalcogens
 - 15 Pnictogens
 - 16 Chalcogens
 - 17 Halogens
 - 18 Helium group
- *Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

87 La 138.905 Lanthanum	88 Ce 140.12 Cerium	89 Pr 140.908 Praseodymium	90 Nd 144.24 Neodymium	91 Pm 144.913 Promethium	92 Sm 150.36 Samarium	93 Eu 151.96 Europium	94 Gd 157.25 Gadolinium	95 Tb 158.93 Terbium	96 Dy 162.50 Dysprosium	97 Ho 164.93 Holmium	98 Er 167.26 Erbium	99 Tm 168.93 Thulium	100 Yb 173.05 Ytterbium
80 Ac 227.028 Actinium	81 Th 232.038 Thorium	82 Pa 231.036 Protactinium	83 U 238.029 Uranium	84 Np 237.048 Neptunium	85 Pu 244.064 Plutonium	86 Am 243.061 Americium	87 Cm 247.070 Curium	88 Bk 247.070 Berkelium	89 Cf 251.08 Californium	90 Es 252.083 Einsteinium	91 Fm 257.095 Fermium	92 Md 258.10 Mendelevium	93 No 259.10 Nobelium

Radioactive

Oganesson is likely to be either a metalloid or a post-transition metal

Active metals

Pre-transition
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Th onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Ir onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metalloids
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750-1,000 kJ/mol), and electronegativity values (1.8-2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colourless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

*as liquid P, the most stable form in ambient conditions.

Group 1
The alkali are Li-F, and the alkaline earths are Ca-Ac, and Al-Lu (except Li, Be, B, C, N, O, F, Ne, Na, Mg, Al, Si, P, S, Cl, Ar, K, Ca, Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Ga, Ge, As, Se, Br, Kr, Rb, Sr, Y, Zr, Nb, Mo, Tc, Ru, Rh, Pd, Ag, Cd, In, Sn, Sb, Te, I, Xe, Ba, La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu, Hf, Ta, W, Re, Os, Ir, Pt, Au, Hg, Tl, Pb, Bi, Po, At, Rn, Fr, Ra, Ac, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr, Oganesson).

Covalent nonmetals
O, F, and Cl are, respectively, colourless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically covalent nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colourless, odourless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answer Sheet

Rachel E. Green



Periodic Table Guide Answer Sheet:

Embracing the Melody of Expression: An Psychological Symphony within **Periodic Table Guide Answer Sheet**

In some sort of used by screens and the ceaseless chatter of quick conversation, the melodic beauty and mental symphony produced by the published term often diminish into the background, eclipsed by the persistent noise and disturbances that permeate our lives. However, situated within the pages of **Periodic Table Guide Answer Sheet** a marvelous fictional prize brimming with organic thoughts, lies an immersive symphony waiting to be embraced. Crafted by a wonderful musician of language, this charming masterpiece conducts viewers on an emotional journey, well unraveling the concealed tunes and profound affect resonating within each carefully constructed phrase. Within the depths of the touching review, we shall investigate the book is main harmonies, analyze their enthralling publishing type, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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