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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $-\log 0.100 = 1.000$ ph $\log 0.100 = 1.000$ b titrant volume 12.50 ml

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Volume of NaOH added (mL)	pH
4.5	4.6
4.6	4.4
4.7	5.9
4.8	5.5
4.9	5.9

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