



To convert mass to energy, use $E=mc^2$
 Annihilation - when a particle and its corresponding antiparticle collide and their rest energy is converted into photons

Pair production - when a photon turns into a particle and its corresponding antiparticle. The photon must have at least the same amount of energy as the rest energies of the particles

Excitation
 Exciting more than just electrons through energy to climb their atomic

Alpha particles - Helium nucleus, ${}^4_2\text{He}$
 - Highly ionising, so it is dangerous
 - Relatively penetrating - stopped by a plate of paper or foil in air

Beta - fast moving electron $\rightarrow e^-$
 - Medium penetrating and relatively ionising - stopped by a few cm of aluminium
 - Beta decay - ${}^0_{-1}\text{e}$ or ${}^0_{+1}\text{e}$ is $\frac{1}{1836}$ of a proton



Gamma - high energy - EM ray or photon emitted from a high energy nucleus
 - Highly penetrating - stopped by a few m of concrete

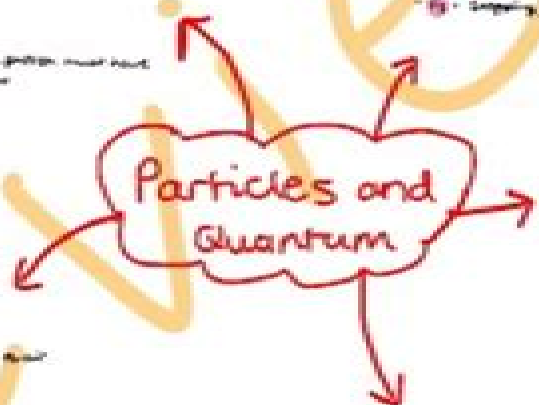
Isotopes - Same element - same number of protons
 - Different relative atomic mass - different number of neutrons
 - e.g. ${}^{12}_6\text{C}$ and ${}^{14}_6\text{C}$

Negative charge - $\frac{Q}{r^2}$ (Coulomb)
 - Charge to mass ratio

Electron volt - charge of an electron when accelerated through a pot of one volt
 - $1.6 \times 10^{-19} \text{ J}$
 - 1 eV

Force	Direction
Gravitational	Downwards
Electromagnetic	Upwards
Nuclear	Outwards

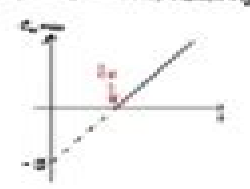
Charge, spin and baryon number must always be conserved
 Strong force is only connected to strong interaction



Photoelectric effect

Use photons of sufficient energy or absorbed by electrons in the surface of a metal, liberating them

- $E_{\text{photon}} = hf$
- $E_{\text{photon}} = \text{photon absorbed} + \text{work function}$ (energy to free electron)
- f - frequency
- h - Planck's constant - minimum energy to liberate electron
- $E_{\text{photon}} = \text{energy of electron} + \text{work function}$
- f_0 - threshold frequency - minimum frequency needed for electrons to be liberated



The photoelectric effect would be particle nature of light due to the discrete threshold frequency and equation $E_k = hf - \phi$ where ϕ is work function. Increasing the intensity of light would increase the number of electrons liberated but it wouldn't affect the kinetic energy of electrons that are emitted per photon

Stopping potential - the pot. needed to stop any electrons from reaching the other side

- $E_{k, \text{max}} = eV_s$
- e - charge of an electron
- V_s - stopping potential

De Broglie

The de Broglie wavelength is the wavelength that a particle can have.
 The wave nature of a particle was proven by diffraction experiments. This is because the electron is diffracting around the crystal lattice.



$$\lambda = \frac{h}{p}$$

λ - de Broglie wavelength
 p - momentum / mass x velocity

Fluorescent tubes

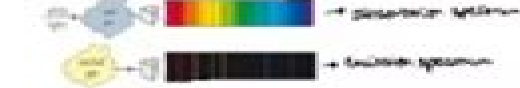


These tubes contain mercury, under a high voltage is applied. This accelerates free electrons to kinetic energy. As the mercury atoms collide with the electrons, they are excited to higher energy levels. When the free electrons return, they emit photons in the visible range. A phosphor coating in the tube absorbs the photons and emits visible light.

Energy levels

Electrons can be excited to higher energy levels. Either by absorbing photons that are equal to the difference in energy levels or by a fast electron collision when transferring the electron's energy.

Ionisation energy - is enough energy to get an electron out of an atom, so will completely remove the electron



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