

CVA Physical Science Unit 6
A6.01.2 Balancing Chemical Equations Assignment



Student Name:
Teacher Name:

Why are we concerned about the amount of each substance present in a chemical reaction? Remember that matter cannot be created nor destroyed during a chemical reaction. This statement is known as the law of conservation of mass.

Chemists use **coefficients** to satisfy the law of conservation of mass. They add coefficients to reactants and products in a chemical equation until there is the same number of each kind of atom on both sides of the equation.

A coefficient in front of a formula indicates the number of molecules of that substance that take part in the chemical reaction. In the reaction $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$, 2 molecules of hydrogen must combine with 1 molecule of oxygen to produce 2 molecules of water. When no coefficient appears before a formula, assume that 1 molecule (or atom) of that substance is present. See Lesson 1 in Module 6 for more information on balancing chemical equations.



Balance the following Chemical Equations. Type the correct **coefficient** in front of the reactants and products so you have the same number of atoms of each element on both sides of the equation. Remember to find the number of atoms, simply multiply the coefficient times the subscript for that atom.

Examples:

Fe is an atom of Iron **2Fe** is 2 atoms of Iron

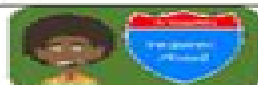
O_2 is a molecule of Oxygen (2 Oxygen atoms bonded together) **2O₂** is 2 Oxygen molecules consisting of 4 Oxygen atoms

(SO₄) is one Sulfate ion (consisting of 1 Sulfur atom and 4 Oxygen atoms), **3(SO₄)** is 3 Sulfate ions (consisting of 3 Sulfur atoms and 12 Oxygen atoms)

Remember, you can only change the coefficient, not the subscript to balance the equation! Save and submit this assignment. Total points 100 Balancing equation 11 is worth 10 extra credit points!

Example: Given - $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ Answer: **2H₂ + O₂ → 2H₂O**

Place your cursor in front of the atom or molecule and type in the coefficient that will balance the chemical equations below:



Place coefficients, where needed in the reactants and or products to balance the following chemical reactions. Remember, you must have the same number of atoms of each element on both sides of the chemical equation! (10 pts. each)

1.



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Table of Contents Physical Science Unit 6 Study Guide

1. Understanding the eBook Physical Science Unit 6 Study Guide
 - The Rise of Digital Reading Physical Science Unit 6 Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Unit 6 Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Unit 6 Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Unit 6 Study Guide
 - Personalized Recommendations
 - Physical Science Unit 6 Study Guide User Reviews and Ratings
 - Physical Science Unit 6 Study Guide and Bestseller Lists
5. Accessing Physical Science Unit 6 Study Guide Free and Paid eBooks
 - Physical Science Unit 6 Study Guide Public Domain eBooks
 - Physical Science Unit 6 Study Guide eBook Subscription Services
 - Physical Science Unit 6 Study Guide Budget-Friendly Options
6. Navigating Physical Science Unit 6 Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Unit 6 Study Guide Compatibility with Devices
 - Physical Science Unit 6 Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Unit 6 Study Guide
 - Highlighting and Note-Taking Physical Science Unit 6 Study Guide
 - Interactive Elements Physical Science Unit 6 Study Guide
8. Staying Engaged with Physical Science Unit 6 Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Physical Science Unit 6 Study Guide
- 9. Balancing eBooks and Physical Books Physical Science Unit 6 Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Unit 6 Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Unit 6 Study Guide
 - Setting Reading Goals Physical Science Unit 6 Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Unit 6 Study Guide
 - Fact-Checking eBook Content of Physical Science Unit 6 Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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