

5 - Periodicity Activity

20 19 18 17 16 15 14 13 12 11 10 9 8 7 6 5 4 3 2 1

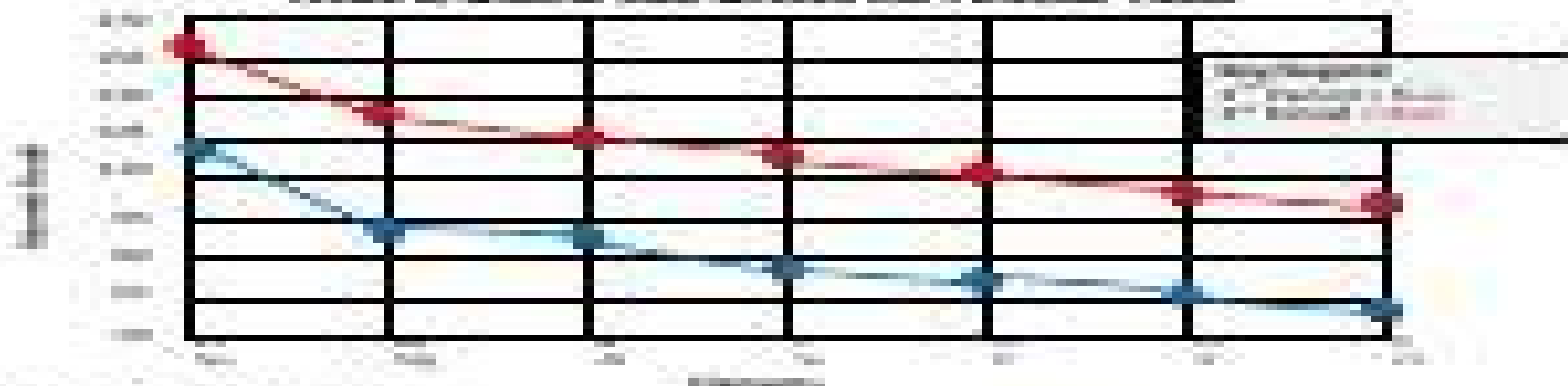
Instructions: Using the graphs above fill the boxes of atomic properties below:

1. Choose which graph best fits the property stated in the graphing sheet of 5 data points (graph above)
2. Is a slight downward trend in this graph Period? If not, fill in the value graph
3. Label the line as acid reactivity increasing
4. Answer all questions...

	1	2	3	4	5	6	7	8
1	1H	2He	3Li	4Be	5B	6C	7N	8O
2	9Na	10Ne	11Mg	12Al	13Si	14P	15S	16Cl
3	17Ar	18Kr	19K	20Ca	21Sc	22Ti	23V	24Cr
4	25Mn	26Fe	27Co	28Ni	29Cu	30Zn	31Ga	32Ge
5	33As	34Se	35Br	36Kr	37Rb	38Sr	39Y	40Zr
6	41Nb	42Mo	43Tc	44Ru	45Rh	46Pd	47Ag	48Cd
7	49In	50Sn	51Sb	52Te	53I	54Xe	55Ba	56La

Now label the trends above in the following way:

Trend in Atomic Size across the Periodic Table



Illustrated Trend Equations:

1. Following the trend in atomic size across the periodic table: decreases
2. This shows two lines that are opposite periodic properties. They are all decreasing (looking downwards) across the periodic table.

Worked Trend Equations:

1. Periodic trend in electronegativity: increases (e.g. Na, Cl, Br, I). It has an increasing movement as you go across to the right of the periodic table: increases
2. This shows two lines that are opposite periodic properties (e.g. size and electronegativity) looking from downwards (downwards)

Periodicity Activity Okstate Answer Key

A Gutmann



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