

Periodic Trends Worksheet

Using the periodic table and your knowledge of periodic trends, answer the following questions:

1 Circle the best answer that has the:

- | | | | | | | |
|---|----------------------------|----|----|----|----|----|
| a | smallest electron affinity | N | P | As | Sb | Bi |
| b | largest electronegativity | Ga | Ge | As | Sc | |
| c | largest ionization energy | O | S | Sc | Te | Po |
| d | lowest ionization energy | O | S | Se | Te | Po |
| e | smallest atomic radius | F | Cl | Br | I | At |
| f | largest atomic radius | Li | Na | K | Rb | Cs |

2 Low ionization energy is a characteristic of a _____

3 What does atomic radius measure? _____

4 Describe the change in the atomic size across a period and down a group.

5 Which element has the largest atomic radius? _____

6 Explain the main difference between metals and nonmetals.

7 Determine if the listed pairs of elements have similar or different properties.

- | | |
|------------------------|----------------------|
| a Potassium & Rubidium | b Calcium & Barium |
| c Sodium & Chlorine | d Helium & Krypton |
| e Phosphorus & Oxygen | f Lithium & Fluorine |

8 Which element has the largest electronegativity? _____

9 Which group (name) would have the lowest ionization energies?

10 Which group (name) has the highest ionization energies?

Periodic Trends And Ions Review Answer Sheet

AW Rasmussen



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Chemistry James E. Brady, Fred Senese, 2004-02-04 Publisher Description

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Chemistry: The Central Science Theodore L. Brown, H. Eugene LeMay Jr., Bruce E. Bursten, Catherine Murphy, Patrick Woodward, Steven Langford, Darius Sagatys, Adrian George, 2013-10-04 If you think you know the Brown LeMay Bursten Chemistry text think again In response to market request we have created the third Australian edition of the US bestseller Chemistry The Central Science An extensive revision has taken this text to new heights Triple checked for scientific accuracy and consistency this edition is a more seamless and cohesive product yet retains the clarity innovative pedagogy functional problem solving and visuals of the previous version All artwork and images are now consistent in quality across the entire text And with a more traditional and logical organisation of the Organic Chemistry content this comprehensive text is the source of all the information and practice problems students are likely to need for conceptual understanding development of problem solving skills reference and test preparation Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Handbook on the Physics and Chemistry of Rare Earths**, 2014-07-10 The Handbook on the Physics and Chemistry of Rare Earths is an ongoing series covering all aspects of rare earth science chemistry life sciences materials science and physics The main emphasis of the Handbook is on rare earth elements Sc Y and the lanthanides La through Lu but information is also included whenever relevant on the closely related actinide elements The individual chapters are comprehensive broad up to date critical reviews written by highly experienced invited experts The series which was started in 1978 by Professor Karl A Gschneidner Jr combines and integrates both the fundamentals and applications of these elements and now publishes two volumes a year Individual chapters are comprehensive broad critical reviews Contributions are written by highly experienced invited experts Up to date overviews of developments in the field **Journal of Research of the National Bureau of Standards** United States. National Bureau of Standards, 1988

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