

Assignment

The Periodic Law

Section Quiz: History of the Periodic Table

As Morganstern provided, enter the name of the person/ people that best completed each statement or label answers as appropriate.

1. In developing his periodic table, Mendeleev's observations were
A. based on atomic number.
B. based on valence electrons.
C. based on atomic weight.
D. based on atomic mass.
E. periodicity.
2. Mendeleev's periodic table distinguished all elements in order of increasing atomic mass because he grouped together elements with
A. similar properties.
B. similar structures.
C. similar atomic numbers.
D. similar atomic masses.
E. similar physical properties.
3. Mendeleev predicted that the gaps in his periodic table represented
A. ions.
B. unknown elements.
C. possible isotopes.
D. noble gases.
E. unknown elements.
4. The periodic table is based on periodicity because the repeating
A. atomic number was
B. atomic weight.
C. atomic mass.
D. atomic number.
E. periodicity.
5. The periodic law describes the physical and chemical properties of elements and predicts descriptions of their atoms.
A. Mendeleev
B. Moseley
C. Bohr
D. Rutherford

Periodic Law Section Quiz Answers

Liyong Dong



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