

**CHAPTER 5 REVIEW***The Periodic Law***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. c In the modern periodic table, elements are ordered  
(a) according to decreasing atomic mass.  
(b) according to Mendeleev's original design.  
(c) according to increasing atomic number.  
(d) based on when they were discovered.
2. d Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing  
(a) density. (c) atomic number.  
(b) reactivity. (d) atomic mass.
3. b The modern periodic law states that  
(a) no two electrons with the same spin can be found in the same place in an atom.  
(b) the physical and chemical properties of an element are functions of its atomic number.  
(c) electrons exhibit properties of both particles and waves.  
(d) the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.
4. c The discovery of the noble gases changed Mendeleev's periodic table by adding a new  
(a) period. (c) group.  
(b) series. (d) level.
5. d The most distinctive property of the noble gases is that they are  
(a) metallic. (c) metalloid.  
(b) radioactive. (d) largely unreactive.
6. c Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of  
(a) 4. (c) 11.  
(b) 10. (d) 18.
7. An isotope of fluorine has a mass number of 19 and an atomic number of 9.  
9 a. How many protons are in this atom?  
10 b. How many neutrons are in this atom?  
 ${}^{19}_{9}\text{F}$  c. What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

# Periodic Law Section Quiz Answer Key

**N Colangelo**



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## **Table of Contents Periodic Law Section Quiz Answer Key**

1. Understanding the eBook Periodic Law Section Quiz Answer Key
  - The Rise of Digital Reading Periodic Law Section Quiz Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Law Section Quiz Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Law Section Quiz Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Law Section Quiz Answer Key
  - Personalized Recommendations
  - Periodic Law Section Quiz Answer Key User Reviews and Ratings
  - Periodic Law Section Quiz Answer Key and Bestseller Lists
5. Accessing Periodic Law Section Quiz Answer Key Free and Paid eBooks

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- Periodic Law Section Quiz Answer Key Budget-Friendly Options
- 6. Navigating Periodic Law Section Quiz Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodic Law Section Quiz Answer Key Compatibility with Devices
  - Periodic Law Section Quiz Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Law Section Quiz Answer Key
  - Highlighting and Note-Taking Periodic Law Section Quiz Answer Key
  - Interactive Elements Periodic Law Section Quiz Answer Key
- 8. Staying Engaged with Periodic Law Section Quiz Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Law Section Quiz Answer Key
- 9. Balancing eBooks and Physical Books Periodic Law Section Quiz Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodic Law Section Quiz Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Law Section Quiz Answer Key
  - Setting Reading Goals Periodic Law Section Quiz Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Law Section Quiz Answer Key
  - Fact-Checking eBook Content of Periodic Law Section Quiz Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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