

# Periodic Table Basics

**Step 1: Complete the squares for each element by adding the atomic number, name, and atomic mass.**

Write the atomic number at the top of the square.

Write the element's name under the symbol.

Write the atomic mass at the bottom of the square.

**Step 2: Determine the number of protons, neutrons, and electrons in each element.**

**Step 3: Create a Bohr diagram for each element.**

**Step 4: Draw the Lewis Structure for each element.**

The card for Boron (B) contains the following information:

- Atomic number: 5
- Symbol: B
- Name: Boron
- Atomic mass: 10.81
- Protons (P): 5
- Neutrons (N): 6
- Electrons (E): 5
- Bohr Diagram: A diagram showing a nucleus with 5 protons and 6 neutrons, and 5 electrons in two shells (2 in the first, 3 in the second).
- Lewis structure: A diagram showing the Boron atom with 5 valence electrons represented by dots around the symbol B.

**Step 5: Use the following colors to shade in the square for each element. You should ONLY color in the small square in the upper left-hand corner and not the entire card.**

Green = Li & Na

Pink = O & S

Blue = Be & Mg

Purple = F & Cl

Orange = B & Al

Red = C & Si

Tan = N & P

Yellow = He, Ne, & Ar

**Step 6: Cut the cards apart and arrange according to atomic number in the pattern shown below. Once you have the cards arranged in the correct order, glue them to a large sheet of construction paper.**



**Step 7: Answer the questions on the back of this worksheet using the information on your Periodic Table.**

# Periodic Table Basics Answers

**LP Steffe**



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## **Table of Contents Periodic Table Basics Answers**

1. Understanding the eBook Periodic Table Basics Answers
  - The Rise of Digital Reading Periodic Table Basics Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Basics Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Table Basics Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Basics Answers

- Personalized Recommendations
  - Periodic Table Basics Answers User Reviews and Ratings
  - Periodic Table Basics Answers and Bestseller Lists
5. Accessing Periodic Table Basics Answers Free and Paid eBooks
    - Periodic Table Basics Answers Public Domain eBooks
    - Periodic Table Basics Answers eBook Subscription Services
    - Periodic Table Basics Answers Budget-Friendly Options
  6. Navigating Periodic Table Basics Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Periodic Table Basics Answers Compatibility with Devices
    - Periodic Table Basics Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Periodic Table Basics Answers
    - Highlighting and Note-Taking Periodic Table Basics Answers
    - Interactive Elements Periodic Table Basics Answers
  8. Staying Engaged with Periodic Table Basics Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Periodic Table Basics Answers
  9. Balancing eBooks and Physical Books Periodic Table Basics Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Periodic Table Basics Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Periodic Table Basics Answers
    - Setting Reading Goals Periodic Table Basics Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Periodic Table Basics Answers

- Fact-Checking eBook Content of Periodic Table Basics Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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