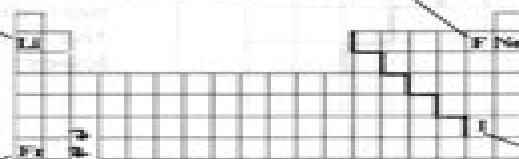


Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

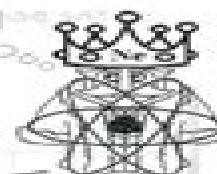
For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

Lithium
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



Fluorine
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (stoutest electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.

I am so perfect... I have a full electron shell. I don't need those other elements... I hope they just leave me alone!



Neon
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

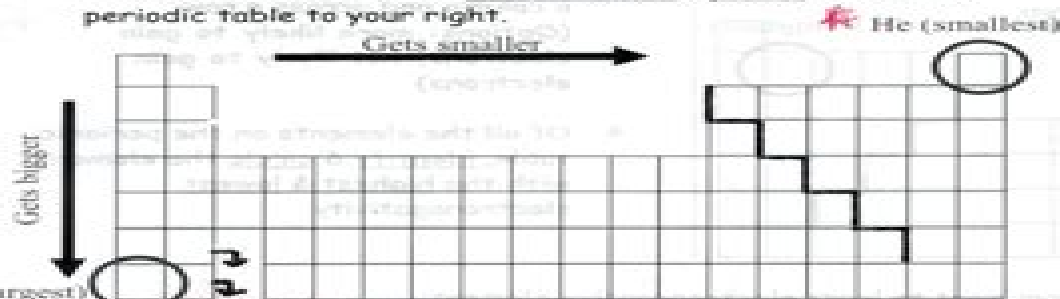
Francium
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

Iodine
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

Part 2: Use your textbook to complete the following:

Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

Effective nuclear charge, Z_{eff} , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.

not a good enough answer
4AP Answer

Periodic Table Answers Pogil

Lauren Gardner

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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1. Understanding the eBook Periodic Table Answers Pogil
 - The Rise of Digital Reading Periodic Table Answers Pogil
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Answers Pogil
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Answers Pogil
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Answers Pogil
 - Personalized Recommendations
 - Periodic Table Answers Pogil User Reviews and Ratings
 - Periodic Table Answers Pogil and Bestseller Lists
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- Periodic Table Answers Pogil Public Domain eBooks
- Periodic Table Answers Pogil eBook Subscription Services
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 - Periodic Table Answers Pogil Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Periodic Table Answers Pogil
 - Highlighting and Note-Taking Periodic Table Answers Pogil
 - Interactive Elements Periodic Table Answers Pogil
- 8. Staying Engaged with Periodic Table Answers Pogil
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Answers Pogil
- 9. Balancing eBooks and Physical Books Periodic Table Answers Pogil
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Answers Pogil
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Table Answers Pogil
 - Setting Reading Goals Periodic Table Answers Pogil
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Table Answers Pogil
 - Fact-Checking eBook Content of Periodic Table Answers Pogil
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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