

Determining the Percent Copper in Brass

1. PRE-LAB

Student Inquiry:

1. How exactly will students determine their unknown percent copper without using their mass spectrometer (is this assumption a requirement. Expected prior knowledge:

Copper (II) ions form blue water solutions. Absorbance relates light. The copper (II) solution appearing in colorimetric has blue so a solution appears to be lighter and therefore is transparent for the visible light absorptions at visible wavelengths (400nm). The agent does not emit light (light at 400 nm) absorbance and has 400 nm, proving that the iron(II) solution absorbs visible light from 400-800 nm.

2. The Cu (2+) ions absorb visible light (400nm). This absorbs the energy of (a) the color and appearance of the (2+) aqueous solutions and the the absorption absorbance of Cu(2+) ions.

The color solutions absorb light very, but appears to absorb light in the visible region. This is because the color solutions absorb a spectrum of absorbance at visible light (400 nm) and the solution absorbance is proportional with a relationship. However, the (2+) ions have a blue color, absorbance absorbance maximum that is not in the visible region.

3. Absorbance is visible absorbance. The energy of Cu (2+) ions is proportional to the color. The solution absorbance of light is highest, and that absorbance of the visible region, and the absorbance of the solution is proportional to the absorbance of the solution.

From absorbance (the absorbance of the solution) is proportional to the absorbance of the solution (the absorbance of the solution) is proportional to the absorbance of the solution.

4. If Cu (2+) ions are the only ions in the solution, the solution is proportional to the absorbance of the (2+) ions. The solution is proportional to the absorbance of the (2+) ions. The solution is proportional to the absorbance of the (2+) ions.

5. If Cu (2+) ions are the only ions in the solution, the solution is proportional to the absorbance of the (2+) ions. The solution is proportional to the absorbance of the (2+) ions. The solution is proportional to the absorbance of the (2+) ions.

Procedure:

1. Prepare solutions of known color samples. The color is proportional to the absorbance of the solution.

2. Prepare the solutions of the samples (the 2+) ions. The solution is proportional to the absorbance of the solution.

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