

1. \_\_\_\_\_ **Ambident bases** \_\_\_\_\_, tend to release hydroxide ions ( $\text{OH}^-$ ) in solution.
2. \_\_\_\_\_ **Bronsted-Lowry acids** \_\_\_\_\_ are substances that donate a proton.
3. \_\_\_\_\_ **Lewis acids** \_\_\_\_\_ are substances that can accept a pair of electrons.
4. The \_\_\_\_\_ **pOH** \_\_\_\_\_ scale describes the hydroxide ion concentration.
5. Acidic solutions have a high concentration of \_\_\_\_\_  **$\text{H}_3\text{O}^+$**  \_\_\_\_\_ ions.
6. Solutions that have a pH above 7 are considered \_\_\_\_\_ **basic** \_\_\_\_\_.
7. A(n) \_\_\_\_\_ **amphoteric** \_\_\_\_\_ substance can act as both a Bronsted-Lowry acid and a Bronsted-Lowry base.
8. \_\_\_\_\_ **Buffers** \_\_\_\_\_ are solutions that resist pH changes.

**Instructions:** For each of the following questions, select the letter that best answers each question based on class notes and discussion. (1.5 points each)

- \_\_\_\_\_ 9. Which of the following is a property of acids?
- A. Have a bitter taste
  - B. Have to turn blue litmus paper red**
  - C. Have to turn red litmus paper blue
  - D. Have a slippery feeling
- \_\_\_\_\_ 10. What is neutralization reactions in aqueous solutions produce?
- A. A conjugate acid and conjugate base
  - B. A weak acid and a strong base
  - C. A salt and water**
  - D. A salt and a conjugate acid
- \_\_\_\_\_ 11. The scientist(s) who classified acids and bases by the ions they produce in solution were
- A. Sarrusent
  - B. Lewis
  - C. Bronsted and Lowry
  - D. Arrhenius**
- \_\_\_\_\_ 12. Using the Bronsted-Lowry model, strong acids

# Reading Essentials Acids Bases And Salts Answer Key

**Julia Schneider**



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