

## 12

## STOICHIOMETRY

**SECTION 12.1 THE ARITHMETIC OF EQUATIONS** (pages 353–358)

*This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.*

**► Using Everyday Equations** (pages 353–355)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You can use the balanced equation.

2. Quantity usually means the amount of a substance expressed in grams or moles.

3. A bookcase is to be built from 3 shelves (Sh), 2 side boards (Sb), 1 top (T), 1 base (B), and 4 legs (L). Write a “balanced equation” for the construction of this bookcase.

$3\text{Sh} + 2\text{Sb} + \text{T} + \text{B} + 4\text{L} = \text{Sh}_3\text{Sb}_2\text{TBL}_4$

**► Using Balanced Chemical Equations** (page 354)

4. Is the following sentence true or false? Stoichiometry is the calculation of quantities in chemical reactions. true

5. Calculations using balanced equations are called stoichiometric calculations.

**► Interpreting Chemical Equations** (pages 356–357)

6. From what elements is ammonia produced? How is it used?

Ammonia molecules are composed of nitrogen and hydrogen; it is used as a fertilizer.

7. Circle the letter of the term that tells what kind of information you CANNOT get from a chemical equation.

a. moles

d. volume

b. mass

e. number of particles

☒ c. size of particles

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**L Towne**



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## **Table of Contents Reading And Study Workbook Chapter 12 Page 145**

1. Understanding the eBook Reading And Study Workbook Chapter 12 Page 145
  - The Rise of Digital Reading Reading And Study Workbook Chapter 12 Page 145
  - Advantages of eBooks Over Traditional Books
2. Identifying Reading And Study Workbook Chapter 12 Page 145
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reading And Study Workbook Chapter 12 Page 145
  - User-Friendly Interface
4. Exploring eBook Recommendations from Reading And Study Workbook Chapter 12 Page 145
  - Personalized Recommendations
  - Reading And Study Workbook Chapter 12 Page 145 User Reviews and Ratings
  - Reading And Study Workbook Chapter 12 Page 145 and Bestseller Lists
5. Accessing Reading And Study Workbook Chapter 12 Page 145 Free and Paid eBooks
  - Reading And Study Workbook Chapter 12 Page 145 Public Domain eBooks
  - Reading And Study Workbook Chapter 12 Page 145 eBook Subscription Services
  - Reading And Study Workbook Chapter 12 Page 145 Budget-Friendly Options
6. Navigating Reading And Study Workbook Chapter 12 Page 145 eBook Formats
  - ePub, PDF, MOBI, and More
  - Reading And Study Workbook Chapter 12 Page 145 Compatibility with Devices
  - Reading And Study Workbook Chapter 12 Page 145 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Reading And Study Workbook Chapter 12 Page 145
  - Highlighting and Note-Taking Reading And Study Workbook Chapter 12 Page 145
  - Interactive Elements Reading And Study Workbook Chapter 12 Page 145
8. Staying Engaged with Reading And Study Workbook Chapter 12 Page 145

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Reading And Study Workbook Chapter 12 Page 145
- 9. Balancing eBooks and Physical Books Reading And Study Workbook Chapter 12 Page 145
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Reading And Study Workbook Chapter 12 Page 145
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Reading And Study Workbook Chapter 12 Page 145
  - Setting Reading Goals Reading And Study Workbook Chapter 12 Page 145
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reading And Study Workbook Chapter 12 Page 145
  - Fact-Checking eBook Content of Reading And Study Workbook Chapter 12 Page 145
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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