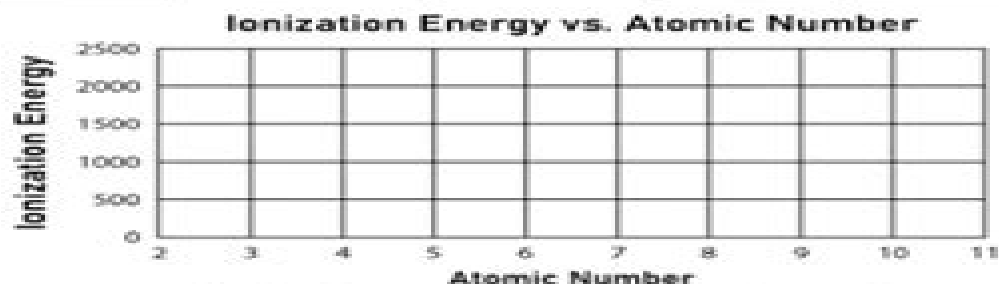


Graphing Periodic Trends Worksheet

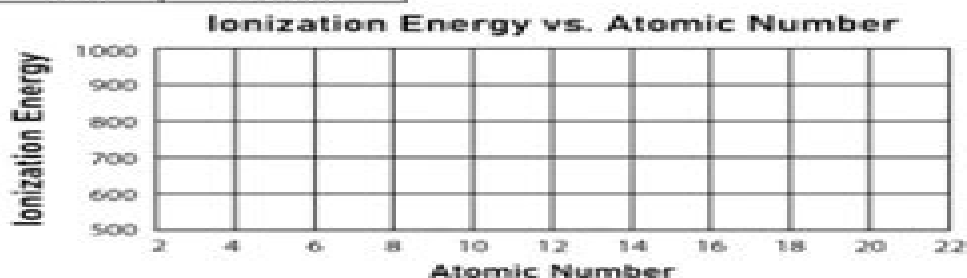
- Using the ionization energy (IE) values of the elements in Period 2 listed below, make a line graph of IE vs. the atomic number.

| Period 2 | | | | | |
|----------|---------------|-------------|---------|---------------|-------------|
| Element | Atomic Number | IE (kJ/mol) | Element | Atomic Number | IE (kJ/mol) |
| Li | 3 | 519 | Li | 3 | 519 |
| Be | 4 | 900 | Be | 4 | 900 |
| B | 5 | 799 | B | 5 | 799 |
| C | 6 | 1088 | C | 6 | 1088 |



- What trend do you notice for the ionization energies in Period 2?
 - Explain why you observe this trend.
- Make a scatter plot of ionization energy vs. atomic number for the first three atoms in Group 2A listed below.

| Group 2A | | |
|----------|---------------|-------------|
| Element | Atomic Number | IE (kJ/mol) |
| Be | 4 | 900 |
| Mg | 12 | 736 |
| Ca | 20 | 590 |



Reading Periodic Trend Graphing Questions

L Towner

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