

Unit 11: Chapter 7 - Chemical Reactions Review Answer Key

Define:

1. **Reactant:** *A substance that undergoes change in a chemical reaction.*
2. **Product:** *New substances formed as a result of a chemical reaction.*
3. **Catalyst:** *A substance that affects the rate of a chemical reaction without being used up in the reaction.*
4. **Coefficient:** *Numbers that appear before a formula in a chemical equation to show the relative proportions of each reactant and product.*
5. **Subscript:** *A number written lower in a chemical formula indicating the amount of atoms present of an element (this number is not to be changed once written in chemical formula).*
6. **Law of Conservation of Mass:** *Matter cannot be created nor destroyed; only changed (atoms can be rearranged).*
7. **Synthesis reaction (include an example):** *A chemical reaction in which two or more substances react to form a single substance.* $(A + B \rightarrow AB)$
8. **Decomposition reaction (include an example):** *A chemical reaction in which a compound breaks down into two or more simpler substances.* $(AB \rightarrow A + B)$
9. **Single replacement (include an example):** *A chemical reaction in which one element takes the place of another element in a compound.* $(AB + C \rightarrow AC + B)$
10. **Double replacement (include an example):** *A chemical reaction in which two compounds exchange positive ions and form two new compounds.* $(AB + CD \rightarrow AD + CB)$
11. **Combustion (include an example):** *A chemical reaction in which a substance reacts rapidly with oxygen, often producing heat and light.* $(C_2H_6 + O_2 \rightarrow CO_2 + H_2O)$
12. **Endothermic reaction:** *A chemical reaction that absorbs energy from its surroundings.*
13. **Exothermic reaction:** *A chemical reaction that releases energy to its surroundings.*
14. **Chemical equation (include an example):** *A representation of a chemical reaction in which the reactants and products are expressed as formulas.*
15. **A balanced equation** is an equation in which the number of atoms on the reactants' side equals the number of atoms on the products' side of the chemical equation.
16. **Equilibrium:** *A state in which the forward and reverse paths of a physical or chemical change take place at the same rate.*

Reactions Review Answer

Arshad Iqbal

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

Reactions Review Answer:

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Table of Contents Reactions Review Answer

1. Understanding the eBook Reactions Review Answer
 - The Rise of Digital Reading Reactions Review Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Reactions Review Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reactions Review Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reactions Review Answer
 - Personalized Recommendations
 - Reactions Review Answer User Reviews and Ratings
 - Reactions Review Answer and Bestseller Lists
5. Accessing Reactions Review Answer Free and Paid eBooks
 - Reactions Review Answer Public Domain eBooks
 - Reactions Review Answer eBook Subscription Services
 - Reactions Review Answer Budget-Friendly Options

6. Navigating Reactions Review Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Reactions Review Answer Compatibility with Devices
 - Reactions Review Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reactions Review Answer
 - Highlighting and Note-Taking Reactions Review Answer
 - Interactive Elements Reactions Review Answer
8. Staying Engaged with Reactions Review Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reactions Review Answer
9. Balancing eBooks and Physical Books Reactions Review Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reactions Review Answer
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reactions Review Answer
 - Setting Reading Goals Reactions Review Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reactions Review Answer
 - Fact-Checking eBook Content of Reactions Review Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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