

Long answer type

1. i) The rate law is given by, $r = [A]^{1/2}[B]^2$. What is the order of this reaction.

ii) A first order reaction is found to have a rate constant $k = 5.5 \times 10^{-14} \text{ s}^{-1}$. Find the Half life of the reaction. iii) Write two differences between order and molecularity.

2. i) A reaction is 50% complete in 20 minutes at 300K and the same reaction again complete in 5 minutes at 350K. Calculate the activation energy if it is first order reaction [$R = 8.314 \text{ JK}^{-1}\text{Mole}^{-1}$; $\log 4 = 0.602$]. ii) Define rate constant and order of a reaction

3. i) A first order reaction has a rate constant of 0.0051 min^{-1} . If we begin with 0.10M concentration of the reactant, what concentration of the reactant will be left after three hour?

ii) A reaction is second order with respect to a reactant. How is its rate effected when concentration of the reactant is

i) doubled ii) reduced to half

4. i) Show graphically how the rate of a 1st order reaction with only one reactant depends upon the concentration of the reactant

ii) Decomposition of phosphine (PH_3) proceeds according to the following equation
 $4\text{PH}_3(\text{g}) \rightarrow \text{P}_4(\text{g}) + 6\text{H}_2(\text{g})$. It is found that the reaction follows the following rate equation $\text{Rate} = k [\text{PH}_3]$
The half life of $[\text{PH}_3]$ is 37.9s at 1200°C .

i) How much time is required for 3/4th of phosphine to decompose.

ii) What fraction of original sample of PH_3 remains behind after 1 minute. Determine the order of the reaction with respect to A and with respect to B and overall order of the reaction.

5(a) A reaction is second order in A and first order in

B. (i). Write the differential rate equation.

(ii). How is the rate affected on increasing the concentrations of A three times?

(iii). How is the rate affected when the concentrations of both A and B are doubled?

(b). A first order reaction takes 40 minutes for 30% decomposition. Calculate Expt. No [A] mole/l [B] mole/l Rate of formation of D mole/l/min
1 0.1 0.1 5.0×10^{-3} 2 0.3 0.2 6.0×10^{-2} 3 0.3 0.4 2.4×10^{-1} 4 0.4 0.1 2.0×10^{-2} $t_{1/2}$ for this reaction. (Given $\log 1.428 = 0.1548$)

Reaction Rate Review Answers

C Cleary



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