

Worksheet 1-1

- Measuring Reaction Rates

Name - _____
 Due Date - _____

Answer the questions in the space provided. You must show all of your work to receive full marks. All answers MUST be rounded to the correct number of Sig. Figures.

1. A chemist wishes to determine the rate of reaction of zinc with hydrochloric acid. The equation for the reaction is:



A piece of zinc is dropped into 1.00 L of 0.100 M HCl and the following data were obtained:

Time	Mass of Zinc
0 s	0.016 g
4 s	0.014 g
8 s	0.012 g
12 s	0.010 g
16 s	0.008 g
20 s	0.006 g

$$\begin{array}{r} 0.016\text{g} \\ - 0.006\text{g} \\ \hline 0.010\text{g} \end{array}$$

- a) Calculate the **Rate of Reaction** in grams of Zn consumed per second.

$$\text{rate} = \frac{\Delta g}{\Delta \text{time}} = \frac{0.010\text{g}}{20\text{sec}}$$

Answer 5.0×10^{-4}

- b) Calculate the **Rate of Reaction** in moles of Zn consumed per second.

$$5.0 \times 10^{-4} \frac{\text{g}}{\text{sec}} \times \frac{1\text{mol}}{65.4\text{g}} =$$

Answer $7.6 \times 10^{-6} \frac{\text{mol}}{\text{sec}}$

- c) Write out the complete ionic equation for the reaction.



- d) What will happen to the $[\text{H}^{+}]$ as the reaction proceeds? decrease

- e) What will happen to the $[\text{Cl}^{-}]$ as the reaction proceeds? nothing, it is a spectator

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