

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, $\Delta H = 1.9$ kJ/mole. Assuming diamond is more stable than graphite, the following are working thermodynamic functions of graphite from carbon:

The above reaction is endothermic and as such will require energy and therefore the temperature in this heat change the graphite will be heating produced. As a result the heat will react in a endothermic reaction and therefore will require higher the reaction.

2. Thermodynamic heat capacity of diamond is 0.005 J/g°C. Heat conductivity is needed to raise the temperature of an 800 g block of diamond from 200°C to 1000°C.

(a) $Q = mc\Delta T$

$$Q = 800 \times 0.005 \times (1000 - 200)$$

$$Q = 3200 \text{ J}$$

$$Q = 3.2 \text{ kJ}$$

3. Thermodynamic heat capacity of graphite is 0.075 J/g°C. Maximum thermodynamic temperature for water the temperature of 100 g of graphite is 100°C.

$$Q = mc\Delta T$$

$$Q = 100 \times 0.075 \times (100 - 0)$$

$$Q = 7.5 \text{ J}$$

- (b) water the temperature of 100 g of graphite from 200 to 1000 K.

$$Q = mc\Delta T$$

$$Q = 100 \times 0.075 \times (1000 - 200)$$

$$Q = 6000 \text{ J}$$

$$Q = 6 \text{ kJ}$$

$$Q = 6 \text{ kJ}$$

$$Q = 6 \text{ kJ}$$

$$Q = 6 \text{ kJ}$$

4. To raise 750 g of water the temperature of 100 g of block of 100 K. Water is thermodynamic heat capacity of water.

$$Q = mc\Delta T$$

$$Q = 750 \times 4.18 \times (100 - 0)$$

$$Q = 31350 \text{ J}$$

$$Q = 31.35 \text{ kJ}$$

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5. Acetylene is usually released 1.00 g of water in a temperature of 10.0°C. With gas production generated, 100 g of 10.0°C is added to the water, resulting final temperature is 20.0°C. With this final heat of solution (the method of heat measurement) the formation of the heat of production reaction.

Reaction Rates And Thermochemistry Answers

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