

## Polyatomic Ions WS

Ammonium ( $\text{NH}_4$ )<sup>1+</sup>  
 Carbonate ( $\text{CO}_3$ )<sup>2-</sup>  
 Chromate ( $\text{CrO}_4$ )<sup>2-</sup>

Dichromate ( $\text{Cr}_2\text{O}_7$ )<sup>2-</sup>  
 Hydroxide ( $\text{OH}$ )<sup>1-</sup>  
 Nitrate ( $\text{NO}_3$ )<sup>1-</sup>

Nitrite ( $\text{NO}_2$ )<sup>1-</sup>  
 Sulfate ( $\text{SO}_4$ )<sup>2-</sup>  
 Sulfite ( $\text{SO}_3$ )<sup>2-</sup>

Phosphate ( $\text{PO}_4$ )<sup>3-</sup>

Write the **NAME** of each of the following compounds.

1.  $(\text{NH}_4)\text{Cl}$  Ammonium Chloride
2.  $\text{Be}(\text{SO}_4)$  Beryllium Sulfate
3.  $(\text{NH}_4)_3\text{N}$  Ammonium Nitride
4.  $\text{MgCl}_2$  Magnesium Chloride
5.  $\text{NH}_4(\text{NO}_3)$  Ammonium Nitrate
6.  $\text{Sr}_3(\text{PO}_4)_2$  Strontium Phosphate
7.  $\text{Zn}(\text{CrO}_4)_2$  Zinc Dichromate
8.  $\text{K}_2(\text{Cr}_2\text{O}_7)$  Potassium Dichromate

9.  $\text{Ga}(\text{ClO}_3)_3$  Gallium Chlorate
10.  $\text{Cu}(\text{OH})$  Copper I Hydroxide
11.  $(\text{NH}_4)_3(\text{PO}_4)$  Ammonium Phosphate
12.  $\text{Fe}(\text{SO}_4)$  Iron II Sulfate
13.  $\text{Mg}(\text{NO}_2)_2$  Magnesium Nitrite
14.  $(\text{NH}_4)\text{NO}_2$  Ammonium Nitrite
15.  $\text{Na}_2(\text{Cr}_2\text{O}_7)$  Sodium Dichromate
16.  $\text{Na}(\text{OH})$  Sodium Hydroxide

Write the **CHEMICAL FORMULA** for each of the given **NAMES**.....cross charges

17. calcium carbonate       $\text{Ca}^{2+} (\text{CO}_3)^{2-}$       =       $\text{Ca}_2(\text{CO}_3)_2$       =  $\text{CaCO}_3$
18. barium nitrate       $\text{Ba}^{+2} \text{NO}_3^{-1}$       =       $\text{Ba}(\text{NO}_3)$
19. ammonium sulfate       $\text{NH}_4^{+1} \text{SO}_4^{-2}$       =       $(\text{NH}_4)_2 \text{SO}_4$
20. aluminum hydroxide       $\text{Al}^{+3} \text{OH}^{-1}$       =       $\text{Al}(\text{OH})_3$
21. calcium phosphate       $\text{Ca}^{+2} \text{PO}_4^{-3}$       =       $\text{Ca}_3(\text{PO}_4)_2$
22. cesium nitrate       $\text{Cs} \text{NO}_2^{-1}$       =       $\text{CsNO}_3$
23. sodium nitrite       $\text{Na}^{+1} \text{NO}_3^{-1}$       =       $\text{NaNO}_3$
24. calcium sulfate       $\text{Ca}^{+2} \text{SO}_4^{-2}$       =       $\text{CaSO}_4$
25. beryllium sulfate       $\text{Be} \text{SO}_4^{-2}$       =       $\text{BeSO}_4$
26. sodium carbonate       $\text{Na}^{+1} \text{CO}_3^{-2}$       =       $\text{Na}_2\text{CO}_3$
27. magnesium phosphate       $\text{Mg}^{+2} \text{PO}_4^{-3}$       =       $\text{Mg}_3(\text{PO}_4)_2$
28. calcium phosphate       $\text{Ca}^{+2} \text{PO}_4^{-3}$       =       $\text{Ca}_3(\text{PO}_4)_2$

# Polyatomic Ions Ws

**Rachel Sandford**



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web in order to have successful collisions the colliding particles must have both the proper amount of energy and the proper fill in blank draw an energy diagram which shows the shape of the curve you might expect for the reaction in this question the overall reaction is exothermic

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web mar 13 2023 work in groups on these problems you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 for the reaction below if substance a is disappearing at a rate of 1.82 times  $10^{-2} \text{ mol l}^{-1} \text{ s}^{-1}$  at what rate is c appearing 3 a 3 b  $\rightarrow$  5 c 2 d

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 $\text{O}_3 \rightarrow \text{O}_2 + \text{O}$  fast  
 $\text{O} + \text{O}_3 \rightarrow \text{O}_2 + \text{O}_2$  slow  
 a write the overall balanced chemical equation  
 $2 \text{O}_3 \rightarrow 3 \text{O}_2$  b identify any intermediates within the mechanism  
 c what is the order with respect

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 consider the following mechanism o3 o2 o fast o3 o 2 o2 slow a write the overall balanced chemical equation 2 o3 3 o2 b  
 identify any intermediates within the mechanism o c

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web reaction rates google classroom mg s 2 hcl a q mgcl a 2 a q h a 2 g the reaction between mg s and hcl a q is represented  
 by the equation above in a kinetics experiment a 0 080 g sample of mg s ribbon is added to an erlenmeyer flask containing 50  
 ml of 3 0 m hcl a q at 25 c

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web jan 10 2021 studying chemical kinetics determines whether to proceed with a reaction as it measures the rate of a  
 reaction reactions conducted in an industrial facility mix compounds together heating and stirring them for a