

# Review Guides for both the midterm and final exams, including detailed PowerPoint Answer Keys

**Periodic Table Review**

Complete the following table by writing the element symbol and name for each element in the appropriate box.

| Group    | 1  | 2  | 13 | 14 | 15 | 16 | 17 | 18 |
|----------|----|----|----|----|----|----|----|----|
| Period 1 | H  | He |    |    |    |    |    |    |
| Period 2 | Li | Be | B  | C  | N  | O  | F  | Ne |
| Period 3 | Na | Mg | Al | Si | P  | S  | Cl | Ar |
| Period 4 | K  | Ca | Sc | Ti | V  | Cr | Mn | Fe |
| Period 5 | Rb | Sr | Y  | Zr | Nb | Mo | Tc | Ru |
| Period 6 | Cs | Ba | La | Hf | Ta | W  | Re | Os |
| Period 7 | Fr | Ra | Ac | Rf | Db | Sg | Bh | Hs |

**Chemical Reactions Review**

Complete the following table by writing the balanced chemical equation for each reaction.

| Reaction                                                                                 | Balanced Equation                                                          |
|------------------------------------------------------------------------------------------|----------------------------------------------------------------------------|
| 1. Hydrogen gas reacts with oxygen gas to form water.                                    | $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$                 |
| 2. Sodium metal reacts with chlorine gas to form sodium chloride.                        | $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$                        |
| 3. Magnesium metal reacts with sulfuric acid to form magnesium sulfate and hydrogen gas. | $\text{Mg} + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + \text{H}_2$ |
| 4. Iron metal reacts with copper(II) sulfate to form iron(II) sulfate and copper metal.  | $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$          |
| 5. Calcium metal reacts with water to form calcium hydroxide and hydrogen gas.           | $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$ |

**Vocabulary**

1. **Physical Property** - a characteristic that can be observed without changing the substance into a different substance.

2. **Chemical Property** - a characteristic that can only be observed when a substance undergoes a chemical change.

3. **Chemical Change** - a change that results in the formation of a new substance.

4. **Physical Change** - a change that does not result in the formation of a new substance.

5. **Law of Conservation of Mass** - matter is neither created nor destroyed in a chemical reaction.

6. **Thermal Energy** - the energy associated with the motion of particles.

7. **Temperature** - a measure of the average kinetic energy of the particles in a substance.

8. **Heat** - the transfer of thermal energy between two objects.

9. **Endothermic** - a process that absorbs thermal energy.

10. **Exothermic** - a process that releases thermal energy.

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- Students play in pairs or small groups.
- Each student has a set of cards with the element symbol and name.
- The teacher reads the element symbol and name.
- The student who has the element symbol and name on their card wins the round.
- The student who wins the round gets to keep their card.
- The game continues until all cards have been used.

Frequency

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# Poe Midterm Study Guide Answer Key

**RJ Alexander**



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