

Chemistry Final Review Study Guide

1. Know what is the atomic number, atomic mass, protons, neutrons and electrons

Atomic number - is the number of protons in the nucleus of an atom of any element

Atomic mass - is expressed in amu; is the weighted average

Protons - are positively charged particles in the nucleus

Neutron - are neutral, have no electrical charge

Electrons - are negatively charged particles in the nucleus

Atomic number = number of protons

Atomic number - (minus) atomic mass = neutrons

2. Describe the shapes of molecules H_2S , CBr_4 ,

and PCl_3

- H_2S - bent
 - Number of bonded atoms: 2
 - Number of lone pairs: 2
 - Number of electron sets: 4

CBr_4 - tetrahedral

- Number of bonded atoms: 4
- Number of lone pairs: 0
- Number of electron sets: 4

PCl_3 - pyramidal

- Number of bonded atoms: 3
- Number of lone pairs: 1
- Number of electron sets: 4

3. Calculating types of bonding (ionic, covalent, nonpolar covalent)

Subtract the smaller EN value from the larger EN value

- 0 - 0.4 = nonpolar covalent
- > (greater than) 0.4 - 1.7 = polar covalent
- > (greater than) 1.7 = ionic bond

Ionic bonds - is a bond that forms when electrons is being transferred from one form to another form, which results in positive and negative ions

Covalent bond - a bond formed when atoms share electrons

Nonpolar covalent bonds - when the EN is very similar between 2 atoms

Polar covalent - when the EN is different between 2 atoms

4. Non - polar covalent bonds and polar covalent bonds and electron sharing

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Will Crane



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