

Plotting Trends

A Periodic Table Activity

Introduction

Most chemistry textbooks report a wealth of numerical data uniformly periodic trends in the properties of elements. Ionization energies, atomic radii, electronegativity values and electron affinities—all are carefully tabulated and graphed. But what do all the numbers mean? In this cooperative activity, use microscale reaction plates and straws of different lengths to construct three-dimensional bar-type charts of the physical properties of the elements. The resulting charts are visually impressive and clearly reveal the meaning of periodic trends in the properties of elements.

Concepts

- Periodic Table
- Periodic trends
- Electron affinity
- Ionization energy
- Atomic radius
- Electronegativity

Materials

(for 28 students working in groups of 4)

Calculator, at least 1 per student group

Handouts of representative elements data, 28

Index cards, 4 × 6 inches, 7

Periodic Table (for reference), at least 1 per student group

Reaction plates, microscale, 96-well (8 × 12 layout), 7

Rulers, metric, marked in millimeters, at least 1 per student group

Schems, at least 1 per student group

Straws, 350 (50 per student group)

Safety Precautions

Although the materials in this activity are considered nonhazardous, follow all normal laboratory safety guidelines.

Procedure

1. Form a working group with three other students.
2. Obtain a handout of Representative Elements Data, a microscale reaction plate, a metric ruler, schems and 50 plastic straws.
3. Find your assigned physical property on the Representative Elements Data handout for 42 representative (outer group) elements.
4. Find the maximum value of the assigned physical property for the elements listed. *Example:* The maximum value of the density for the elements listed is 11.342 g/cm³ (for lead).

5. Let the length of the straw minus one cm represent this maximum value. This will be the scale for all of the other measurements of that property. *Example:* For a straw that is 19.5 cm long, a straw length of 18.5 cm would represent a density of 11.342 g/cm³. The scale is thus 18.5 cm = 11.342 g/cm³. Round off straw lengths to 0.1 cm (1 mm).
6. Using this "straw" scale as a ratio, calculate the straw length that is needed to represent the assigned property for each element in the list. *Example:* The density of beryllium is 1.848 g/cm³. Solving Equation 1 for the straw length (s) shows that a straw length of 3.0 cm would represent the density of beryllium. Round off all straw lengths to 0.1 cm.

$$\frac{18.5 \text{ (cm)}}{11.342} = \frac{s \text{ (cm)}}{1.848}$$

$$(18.5 \times 1.848) / 11.342 = 3.0 \text{ cm} \quad \text{Equation 1}$$

7. Add 1 cm to the calculated straw length for each element, and cut a straw to that length. *Example:* Cut a straw 4.0 cm long to represent beryllium.
8. Place the straw in the reaction plate according to its position in the periodic table. Remember, the transition elements and lanthanides and actinides are not included in the list of representative elements. *Example:* Beryllium is placed in row 2, column 2.
9. Repeat steps 6–8 for each element in the list.
10. Determine the nature of any periodic relationship or trend that may exist for the assigned property of the elements.
11. Propose an explanation for the observed trend.
12. Create a descriptive card to be displayed with the three-dimensional chart. Include the following information on the descriptive card:
 - Names of group members
 - The assigned physical property of the elements
 - Description of the observed trend
 - Proposed explanation for the observed trend

Plotting Trends A Periodic Table Activity

Annie Termaat, Christopher Talbot



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Teacher's Guide to Using the Next Generation Science Standards With Gifted and Advanced Learners Cheryll M. Adams, Alicia Cotabish, Debbie Dailey, 2021-09-23 A Teacher's Guide to Using the Next Generation Science Standards With Gifted and Advanced Learners provides teachers and administrators with practical examples of ways to build comprehensive coherent and rigorous science learning experiences for gifted and advanced students from kindergarten to high school. It provides an array of examples across the four domains of science: physical sciences, Earth and space sciences, life sciences, and engineering technology and applications of science. Each learning experience indicates the performance expectation addressed and includes a sequence of activities, implementation examples, connections to the CCSS Math and CCSS ELA, and formative assessments. Chapters on specific instructional and management strategies, assessment, and professional development suggestions for implementing the standards within the classroom will be helpful for both teachers and administrators.

Transition Metal Sulphides Th. Weber, H. Prins, R.A. van Santen, 2013-06-29 Hydrotreating catalysis with transition metal sulphides is one of the most important areas of industrial heterogeneous catalysis. The present book deals with the chemical and catalytic aspects of transition metal sulphides, focusing on their use in hydrotreating catalysis. The book's 12 chapters present reviews of solid state coordination and organometallic chemistry, surface science, and spectroscopic studies, quantum chemical calculations, catalytic studies with model and real catalysts, as well as refinery processes. A presentation of state of the art background to pertinent work in the field can be used as an introduction to the chemical and catalytic properties of transition metal sulphides, as well as an advanced level reference. 15 TGT Science Test Papers EMRS Mocktime Publication, EMRS Exam Teachers TGT Science Test Papers 15 Practice Papers Tier 1 Eklayva Model Residential Schools as per Official Exam Pattern and Syllabus Software for Teaching Science Roger Frost, 1998

Chemistry: Molecules, Matter, and Change Media Activities Book Loretta Jones, Carl Hoeger, Peter William Atkins, Regina Schoenfeld-Tacher, 2000-01-15 The Media Activity Book (MAB) for Jones Atkins Chemistry: molecules, matter, and change contains chapters with lists and descriptions of some of the media available as you study the chapter. Each activity begins with a specific textbook reference. Then you are given a time estimate of how long it will take to use the media. An M media icon in the margin of the textbook means that media exists to support that area of text. The media is found in three different places: on the website and on two CDs. TRAC: Trends in Analytical Chemistry Y. Gohshi, J.F.K. Huber, T. Townshend, 2013-09-17

Trends in Analytical Chemistry Volume 12 focuses on the advancements of processes, technologies, automation, and applications of analytical chemistry. The selection first offers information on single cell analysis at the level of a single human erythrocyte and micellar catalysis in reaction rate methods. Topics include analytical strategies, analysis of single erythrocytes, kinetic aspects of micellar catalysis, and micellar kinetic multicomponent determination. The text then takes a look at advances in the field of laser atomic spectroscopy and molecular recognition of sugars, including detection of sugar.

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Theoretical Aspects of Heterogeneous Catalysis John B. Moffat, 2013-03-09 Catalysis involves just about every field of scientific study This means that a multidisciplinary approach is needed in catalytic studies Catalysis involves breaking and forming new bonds and this requires an understanding of either adsorption by bonding to an extended structures or bonding in a coordination sphere Any understanding of catalytic action must necessarily involve an understanding of this bonding Even 200 years ago scientists were aware that a properly treated material such as charcoal could adsorb an enormous quantity of gas In 1812 de Sassaure English translation Annal Philosophy 6 241 1815 proposed that the ability of a material to increase the rate of chemical reaction was due to adsorption of the material in the fine structure of the solid so that the concentrations of the reactants were significantly increased and this increase in concentration led to an increase in reaction rate During the 1800s little advance was made in the understanding of adsorption

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Table of Contents Plotting Trends A Periodic Table Activity

1. Understanding the eBook Plotting Trends A Periodic Table Activity
 - The Rise of Digital Reading Plotting Trends A Periodic Table Activity
 - Advantages of eBooks Over Traditional Books
2. Identifying Plotting Trends A Periodic Table Activity
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Plotting Trends A Periodic Table Activity
 - User-Friendly Interface
4. Exploring eBook Recommendations from Plotting Trends A Periodic Table Activity
 - Personalized Recommendations

- Plotting Trends A Periodic Table Activity User Reviews and Ratings
- Plotting Trends A Periodic Table Activity and Bestseller Lists
- 5. Accessing Plotting Trends A Periodic Table Activity Free and Paid eBooks
 - Plotting Trends A Periodic Table Activity Public Domain eBooks
 - Plotting Trends A Periodic Table Activity eBook Subscription Services
 - Plotting Trends A Periodic Table Activity Budget-Friendly Options
- 6. Navigating Plotting Trends A Periodic Table Activity eBook Formats
 - ePub, PDF, MOBI, and More
 - Plotting Trends A Periodic Table Activity Compatibility with Devices
 - Plotting Trends A Periodic Table Activity Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Plotting Trends A Periodic Table Activity
 - Highlighting and Note-Taking Plotting Trends A Periodic Table Activity
 - Interactive Elements Plotting Trends A Periodic Table Activity
- 8. Staying Engaged with Plotting Trends A Periodic Table Activity
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Plotting Trends A Periodic Table Activity
- 9. Balancing eBooks and Physical Books Plotting Trends A Periodic Table Activity
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Plotting Trends A Periodic Table Activity
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Plotting Trends A Periodic Table Activity
 - Setting Reading Goals Plotting Trends A Periodic Table Activity
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Plotting Trends A Periodic Table Activity
 - Fact-Checking eBook Content of Plotting Trends A Periodic Table Activity

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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