

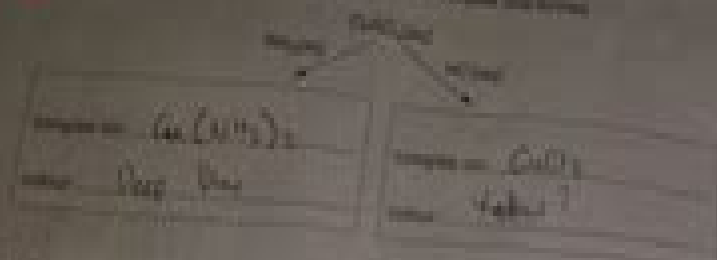
Question 1

Answer all the questions

The reaction is about the synthesis of copper complexes and ligands.

(a) The reaction shows the synthesis of a copper complex.

In the boxes, write the formula and name of the ligands and metal.



(b) Cu^{2+} has two d-orbitals in the d_{xy} and d_{yz} orbitals. The d_{xy} orbital is a d-orbital.

The orbital formula of the d-orbital is shown in Fig. 1.1 below.



Fig. 1.1

(c) What is meant by the term d-orbital?

A d-orbital is a type of atomic orbital that has two lobes, one above and one below the central nucleus.

(a) The reaction is shown below.

The reaction is about the synthesis of a copper complex.



Complex A

Complete the 3D diagram of the other two geometries of A.

Indicate with ticks whether the stereoisomers are cis, trans, optical or a combination of these types.

Stereoisomer														
Type	<table border="1" data-bbox="1244 1023 1404 1170"> <tr> <td>cis</td> <td></td> </tr> <tr> <td>trans</td> <td>✓</td> </tr> <tr> <td>optical</td> <td></td> </tr> </table>	cis		trans	✓	optical		<table border="1" data-bbox="1584 1023 1744 1170"> <tr> <td>cis</td> <td>✓</td> </tr> <tr> <td>trans</td> <td></td> </tr> <tr> <td>optical</td> <td>✓</td> </tr> </table>	cis	✓	trans		optical	✓
cis														
trans	✓													
optical														
cis	✓													
trans														
optical	✓													

(b) What is the empirical formula, including the charge, of the complex in A?



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Hussin A.Rothana



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