

16. The dissociation of water is shown below:



- At 25 °C, the auto product of water, K_w , is 1.0×10^{-14} mol² dm⁻⁶.
At 65 °C, the auto product of water, K_w , is 1.0×10^{-13} mol² dm⁻⁶.

17. Explain what the dissociation of water is an indication of water's chemical nature.

Endothermic.

18. Predict, using a calculation, whether a solution of 1.0 mol dm⁻³ is neutral, acidic or alkaline.

Acidic

Explain by calculating pH.

19. pK_w , pH and pOH are significant notions.

Calculate pK_w at 65 °C.

Give your answer to two decimal places.

$$-\log(1.0 \times 10^{-13}) = 13.00$$

$$pK_w = 13.00$$

20. A solution of hydrochloric acid is mixed with water and the resulting solution is neutral at 25 °C.

Calculate the pH of the resulting solution at 65 °C.

Give your answer to two decimal places.

$$pH = 7.00$$

pH = _____ (1)

(Total: 10)

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