

c. Standard Enthalpy of Vaporization ΔH_{vap}°

→ The enthalpy change for the vaporization of 1 mole of liquid to a gas at the boiling point and standard pressure.

Example:

$$\Delta H_{vap}^\circ = \Delta H_{gas}^\circ - \Delta H_{liq}^\circ \quad (\Delta H_{vap}^\circ = 44.0 \text{ kJ mol}^{-1})$$

d. Standard Enthalpy of Melting ΔH_{fus}°

→ The enthalpy change for melting 1 mol of a solid into a liquid at the melting point and standard pressure.

Example:

$$\Delta H_{fus}^\circ = \Delta H_{liq}^\circ - \Delta H_{sol}^\circ \quad (\Delta H_{fus}^\circ = 6.01 \text{ kJ mol}^{-1})$$

e. Standard Neutralization Enthalpy Change ΔH_{neut}°

→ The enthalpy change in the neutralization of an acid (H⁺) by a base (OH⁻) forming 1 mole of water.

Example:

$$\Delta H_{neut}^\circ = \Delta H_{aq}^\circ(\text{products}) - \Delta H_{aq}^\circ(\text{reactants}) \quad (\Delta H_{neut}^\circ = -57.3 \text{ kJ mol}^{-1})$$

f. Standard Enthalpy of Formation $\Delta H_{f,298}^\circ$

→ The enthalpy change for forming 1 mol of elemental atoms in the gas phase at standard conditions.

Example:

$$\Delta G_f^\circ = \Delta G_{gas}^\circ \quad (\Delta G_{gas}^\circ = 46.2 \text{ kJ mol}^{-1})$$

g. Standard Solution Enthalpy Change ΔH_{sol}°

→ The enthalpy change when 1 mole of a compound is converted into a solution at standard conditions.

Example:

$$\Delta H_{sol}^\circ = \Delta H_{aq}^\circ \quad (\Delta H_{sol}^\circ = -23 \text{ kJ mol}^{-1})$$

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Jessica J Manson



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