

Instructions

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August, 2012

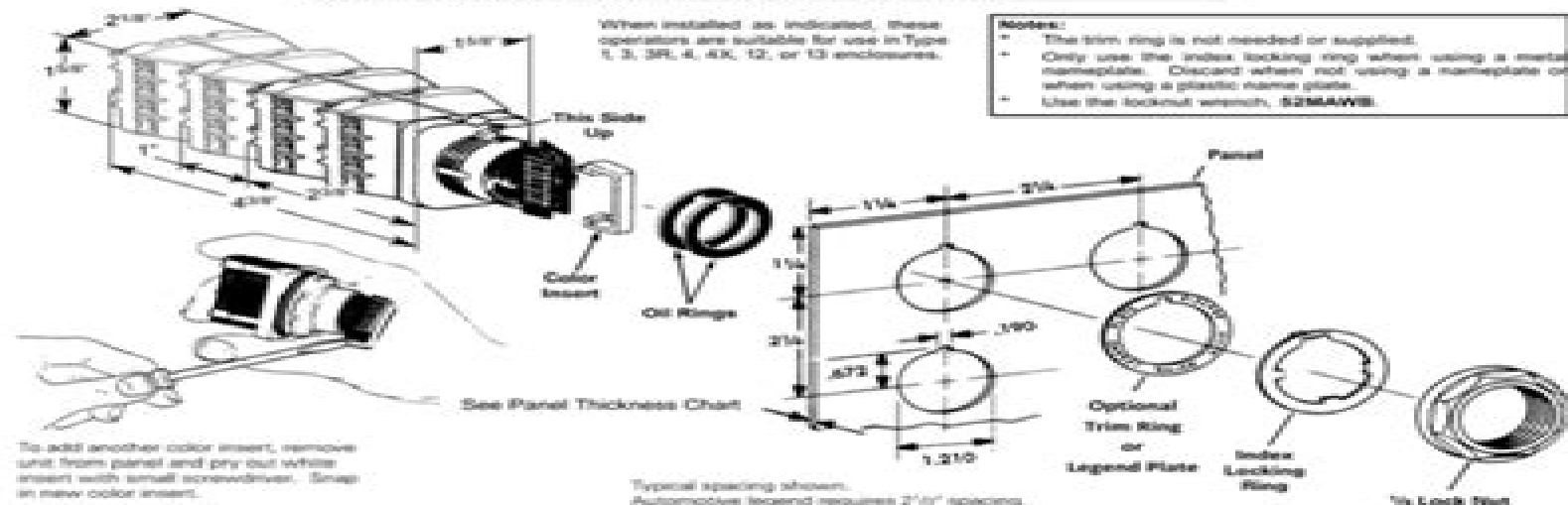
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E87910-A0070-T003-A3-52SX

Selector Operator
Class 52
Cat. No. 52SA2, 52SB2
52SX2 & 52SW2



WARNING

Hazardous voltage.
Can cause death, serious personal injury, or property damage.

To avoid electrical shock, turn off electrical power to controls before performing installation/maintenance.

INSTALLATION

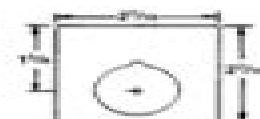
Units are assembled as shown.

- Position and press in color insert if required.
- Assemble oil rings.
- Mount unit on panel as shown in the illustration.
- * Position optional trim ring or legend plate if used.
- * Position index locking ring on unit as shown in illustration.
- Assemble and tighten lock nut to 40 +/- 2 in.-lbs.
* Locking ring not needed with plastic legend plates.

Panel Thickness Chart

Panel Thickness Range (inches)	Gaskets Required
16 ga 0.060 - 0.069	2
15 ga - 11 ga (0.070 - 0.125)	1

LEGEND PLATES



Sea Siemens Specification Guide

Emilie Sanchez



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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a byproduct of the process as iron ions the iron metal dissolves

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web purpose to enhance the understanding of stoichiometry a reaction between iron and copper ii sulfate solution will be conducted this will help you to differentiate limiting and excess reactant in a chemical reaction finally the theoretical and

percent yield of this reaction will be calculated

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web october 3 2017 abstract the techniques used in this lab are quantitative transfer and vacuum filtration with the reaction of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine the product of copper ii sulfate with iron

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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a piece of filter paper taking the filter paper s mass into account add these to the beaker 3 measure 50 0 ml of water in the cylinder and add it to the beaker 4

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web stoichiometry lab the reaction of iron with copper ii sulfate the study of stoichiometry deals with the calculation of quantities in a chemical reaction how much product will be produced how much reactant do you need to make that much product

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web description great lab for use during stoichiometry unit in chemistry reinforces ideas of mole ratios and stoichiometry calculations by measuring the mass of iron that reacts and the mass of copper metal produced students can calculate the ratio of moles of reactant to moles of product this experimental mole ratio from lab can be compared

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away copper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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stirring the solution adding water then removing water adding acetone heating the acetone on low

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