

$$\Delta G^\circ = -2.85 \text{ kJ}$$

ह) मान निकाल  $E^\circ$  की सहा

$$\Delta G^\circ = -2.85 \text{ kJ}$$

$$\log K_c = \frac{-\Delta G^\circ}{2.303RT}$$

$$= \frac{(-2.85)}{2.303 \times 8.314 \times 298} = -0.5014$$

$$\log K_c = -0.5014$$

$$K_c = 0.32$$

5. निम्नलिखित सेले को प्ले X पर सेल स्क्रिप्ट लिखें and

निर्धार



दिए गए हैं

$$E^\circ_{\text{Mg}^{2+}/\text{Mg}} = -2.37 \text{ V}$$

$$E^\circ_{\text{Fe}^{2+}/\text{Fe}} = -0.44 \text{ V}$$

$$E^\circ_{\text{Sn}^{2+}/\text{Sn}} = -0.14 \text{ V}$$

$$E^\circ_{\text{Pb}^{2+}/\text{Pb}} = -0.13 \text{ V}$$

$$E^\circ_{\text{H}^+/\text{H}_2} = 0.00 \text{ V}$$

$$E^\circ_{\text{Co}^{2+}/\text{Co}} = -0.28 \text{ V}$$

ह) दिए गए सेल के लिए सेल स्क्रिप्ट लिखें and

निर्धार करें

(a)  $\text{Zn} \parallel \text{Zn}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(b)  $\text{Fe} \parallel \text{Fe}^{2+} (0.001 \text{ M}) \parallel \text{H}^+ (0.001 \text{ M}) \parallel \text{H}_2 \text{ (g)}$

(c)  $\text{Sn} \parallel \text{Sn}^{2+} (0.001 \text{ M}) \parallel \text{H}^+ (0.001 \text{ M}) \parallel \text{H}_2 \text{ (g)}$

(d)  $\text{Pb} \parallel \text{Pb}^{2+} (0.001 \text{ M}) \parallel \text{H}^+ (0.001 \text{ M}) \parallel \text{H}_2 \text{ (g)}$

(e)  $\text{Zn} \parallel \text{Zn}^{2+} (0.001 \text{ M}) \parallel \text{Fe}^{2+} (0.001 \text{ M}) \parallel \text{Fe}$

(f)  $\text{Fe} \parallel \text{Fe}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(g)  $\text{Sn} \parallel \text{Sn}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(h)  $\text{Pb} \parallel \text{Pb}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(i)  $\text{Zn} \parallel \text{Zn}^{2+} (0.001 \text{ M}) \parallel \text{Fe}^{2+} (0.001 \text{ M}) \parallel \text{Fe}$

(j)  $\text{Fe} \parallel \text{Fe}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(k)  $\text{Sn} \parallel \text{Sn}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(l)  $\text{Pb} \parallel \text{Pb}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(m)  $\text{Zn} \parallel \text{Zn}^{2+} (0.001 \text{ M}) \parallel \text{Fe}^{2+} (0.001 \text{ M}) \parallel \text{Fe}$

(n)  $\text{Fe} \parallel \text{Fe}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(o)  $\text{Sn} \parallel \text{Sn}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

(p)  $\text{Pb} \parallel \text{Pb}^{2+} (0.001 \text{ M}) \parallel \text{Cu}^{2+} (0.001 \text{ M}) \parallel \text{Cu}$

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