

Chapter 12 – Stoichiometry

Note Taking Guide: Episode 801 Stoichiometry

-Study of the **QUANTITY** relationships in a **CHEMICAL REACTION**

-Based on **BALANCED** equations



1. Particles

2 atoms Mg	2 atoms O	2 atoms Mg
		2 atoms O

2 atoms Mg	1 molecule O ₂	2 molecule MgO
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10 atoms Mg	5 molecule O ₂	10 molecule MgO
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2 (6.02 x 10 ²³) atoms Mg	1 (6.02 x 10 ²³) molecule O ₂	2 (6.02 x 10 ²³) molecule MgO
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2. Moles

2 mol Mg	1 mol O ₂	2 mol MgO
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The **COEFFICIENTS** in a **BALANCED EQUATION** give the **MOLE RATIO** for the **SUBSTANCES** involved in the **REACTION**.

Lab: Episode 801 - Can You Find the Mole Ratio?

In this experiment, baking soda (NaHCO₃) reacts with hydrochloric acid (HCl) to produce sodium chloride, carbon dioxide, and water. You will determine the moles of reactant used and moles of product produced using both the lab data and stoichiometry and compare the two.

Data Table:

mass of test tube and baking soda	19.46 g
mass of empty test tube	19.21 g

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RJ Alexander



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