

### Recap/Summary

1. Study of this \_\_\_\_\_ method \_\_\_\_\_ relationships in a \_\_\_\_\_ situation \_\_\_\_\_
2. Research on \_\_\_\_\_ behaviour \_\_\_\_\_ equated to \_\_\_\_\_

2.  $40g + 40g \rightarrow 2 \text{ MgO}$

2. Magnesium burning with oxygen to produce magnesium oxide  
 2. Magnesium burning with oxygen to produce magnesium oxide

The \_\_\_\_\_ in a \_\_\_\_\_ reaction \_\_\_\_\_ give the \_\_\_\_\_ \_\_\_\_\_ from the \_\_\_\_\_ involved in the reaction \_\_\_\_\_

E.g. Problem

When aluminium reacts with oxygen to form aluminium oxide, aluminium reacts to produce



moles of Al

4	Al	3	O <sub>2</sub>
2	Al	2	Al <sub>2</sub> O <sub>3</sub>
2	O <sub>2</sub>	2	Al <sub>2</sub> O <sub>3</sub>

2. If you start with 4 moles of Al, how many moles of Al<sub>2</sub>O<sub>3</sub> will be produced?

2. 2 moles Al (from Al<sub>2</sub>O<sub>3</sub>) and 3 moles O<sub>2</sub> = 2 moles Al<sub>2</sub>O<sub>3</sub>

# Note Taking Guide Episode 801 Stoichiometry Answers

**M Lipman**



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